

Equilibrium Expressions Worksheet

Look at the chemical equation $aA + bB \rightleftharpoons cC + dD$. In this reaction A, B, C, and D are chemical substances, and a, b, c, and d are the coefficients in the chemical equation. Since the equation is reversible, the reactants will not necessarily run to completion—so there won't be a limiting reactant. The forward equation and the backward equation will reach an equilibrium point where the forward equation is using A and B as fast as the backwards equation is making A and B.

If the temperature is held constant, the reaction will always reach equilibrium in such a way such that the equilibrium constant is the same. (At a different temperature the equilibrium constant for the reaction would be a constant but would be a different number.)

Equilibrium constant (K_{eq})

The relationship between the concentration of products and reactants at equilibrium can be expressed by an equilibrium constant K_{eq} (or K). Given the chemical equation in the form $aA + bB \rightleftharpoons cC + dD$, K_{eq} is found by the equation

$$K_{eq} = \frac{[C]^c [D]^d}{[A]^a [B]^b}$$

Note: [] denotes molar concentrations.

But when writing equilibrium expressions, reactants or products that are pure solids or pure liquids do not appear in the expression since no matter how much of it there is its density, and therefore its concentration, is always the same.

1. Write equilibrium expressions for each in the box provided:

- a. $O_3(g) + NO(g) \rightleftharpoons O_2(g) + NO_2(g)$
- b. $2CO(g) + O_2(g) \rightleftharpoons 2CO_2(g)$
- c. $NH_4NO_3(s) \rightleftharpoons N_2O(g) + 2H_2O(l)$
- d. $2H_2O(g) \rightleftharpoons 2H_2(g) + O_2(g)$

a.

b.

c.

d.

Calculations involving K_{eq}

The concentrations of the solutions/gases involved in the reversible reactions **at equilibrium** are known as the equilibrium concentrations. Since K_{eq} represents the system at equilibrium, only the equilibrium concentrations can be used in the equilibrium expression (not the initial concentrations).

2. $PCl_3(g) \rightleftharpoons PCl_5(g) + Cl_2(g)$. What is the equilibrium constant if the equilibrium concentrations are as follows: PCl_5 is 0.0096 mol/L, PCl_3 is 0.0247 mol/L and Cl_2 is 0.0247 mol/L? Ans: 0.064

3. At 1000°C, a 1.00 L container has an equilibrium mixture consisting of 0.102 mol of ammonia, 1.03 mol of nitrogen, and 1.62 mol of hydrogen. Calculate the K_{eq} for the equilibrium system. $N_2(g) + 3H_2 \rightleftharpoons 2NH_3(g)$ Ans: 0.00238

4. At a given temperature, the K_{eq} for the reaction $2HI(g) \rightleftharpoons H_2(g) + I_2(g)$ is 1.40×10^{-2} . If the concentration of both H_2 and I_2 at equilibrium are $2.00 \times 10^{-4}M$, find the concentration of HI. Ans: 0.00169M

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