Explain how you identified the limiting and excess reactant based only on your qualitative observations (not quantitative measurements or calculations). Type your answer into the box. Use complete sentences.

According to observation the initial color of browningsgreegy purple is yellow in color after adding the NaOH, it turns violet. After Changing the color yellow from purple indicates acid is completely consumed so excess of base addition changes the color into violet.

Table C. Calculations to Determine Theoretical Yields

Description	Calculated Value
Volume of NaOH solution	22.20 mL
Initial moles of NaOH	0.002451 moles
Initial moles of H ₃ PO ₄	0.002602 moles
Mass of H ₂ O produced if NaOH runs out first	0.04414 g
Mass of H ₂ O produced if H ₃ PO ₄ runs out first	1.265 g
Theoretical yield of H ₂ O (in grams)	1.265 g
Theoretical yield of Na ₃ PO ₄ (in grams)	0.4267 g

Table D. Identification of Limiting and Excess Reactants from Calculations

Description	ption Reactant Name or Formula	
Limiting Reactant	Phosphoric Acid	
Excess Reactant	Sodium Hydroxide	

Explain how you identified the limiting and excess reactant based on your calculations. Be specific. Type your answer into the box. Use complete sentences.

According to our calculation the reacting equivalent of H₃PO₄ is less than NaOH. Therefore H₂PO₄ is limiting reactant so product formation depends on H₃PO₄ because phosphoric acid is completely consumed.

Formal Lab Report On Limiting Reactant

Dr. Priyanka Gupta Manglik

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Exploring General, Organic, & Biochemistry in the Laboratory William G. O'Neal,2017-02-01 This full color comprehensive affordable manual is appropriate for two semester introductory chemistry courses It is loaded with clearly written exercises critical thinking questions and full color illustrations and photographs providing ample visual support for experiment set up technique and results

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Information Circular ,1988 Nuclear Science Abstracts ,1975-08 Calculating Theoretical Yield Is Not Impossible! Michael S. Leonard, 2025-02-25 Calculating theoretical and percent yield is a fundamental skill for the laboratory This book primarily targets Organic Chemistry Laboratory courses at the high school or college and university level as a supplemental resource to help students master this skill It begins with simple examples from everyday life demonstrates the importance of balancing the equation addresses the role of the mole in these computations discusses different types of liquids considers the role of significant figures and culminates with the planning of syntheses There are suggestions for further reading as well as practice problems and questions to ensure mastery Begins with examples from everyday life that enable students to understand the concepts of theoretical and percent yield before applying those concepts to the laboratory Addresses the necessity of balancing the reaction equation the centrality of the mole in these calculations and the role of significant figures in reporting the answer Explains how to approach the calculations when using neat liquids or solutions The culmination of this text is the use of the same thought processes to plan the amounts of reactants needed for syntheses of desired quantities of product All of the problems in the book include detailed solutions with accompanying text to explain the answers and ancillaries also include suggestions for further reading Mine Drainage and Surface Mine Reclamation, 1988

Multiscale Operational Organic Chemistry John W. Lehman, 2009 This comprehensive laboratory text provides a

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