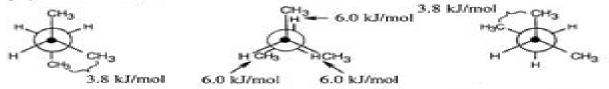
3.42

The energy difference between the two conformations is (11.0 + 6.0 + 4.0) kJ/mol – 3.8 kJ/mol = 17.2 kJ/mol.

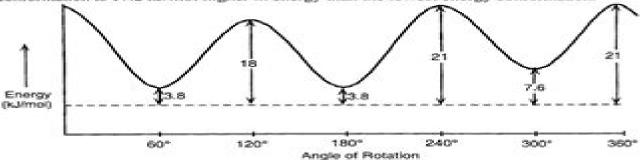
(c) Consider the least stable conformation to be at zero degrees. Keeping the front of the projection unchanged, rotate the back by 60° to obtain each conformation.



at 120°; energy = 3.8 kJ/mol at 120°; energy = 18.0 kJ/mol at 180°; energy = 3.8 kJ/mol



Use the lowest energy conformation as the energy minimum. The highest energy conformation is 17.2 kJ/mol higher in energy than the lowest energy conformation.



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Peter Atkins

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Table of Contents Mcmurry Organic Chemistry Solutions Manual Scribd

- 1. Understanding the eBook Mcmurry Organic Chemistry Solutions Manual Scribd
 - The Rise of Digital Reading Mcmurry Organic Chemistry Solutions Manual Scribd
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Mcmurry Organic Chemistry Solutions Manual Scribd
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Mcmurry Organic Chemistry Solutions Manual Scribd
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Mcmurry Organic Chemistry Solutions Manual Scribd
 - Personalized Recommendations
 - Mcmurry Organic Chemistry Solutions Manual Scribd User Reviews and Ratings
 - Mcmurry Organic Chemistry Solutions Manual Scribd and Bestseller Lists
- 5. Accessing Mcmurry Organic Chemistry Solutions Manual Scribd Free and Paid eBooks
 - Mcmurry Organic Chemistry Solutions Manual Scribd Public Domain eBooks
 - Mcmurry Organic Chemistry Solutions Manual Scribd eBook Subscription Services
 - Mcmurry Organic Chemistry Solutions Manual Scribd Budget-Friendly Options
- 6. Navigating Mcmurry Organic Chemistry Solutions Manual Scribd eBook Formats
 - ePub, PDF, MOBI, and More
 - Mcmurry Organic Chemistry Solutions Manual Scribd Compatibility with Devices
 - Mcmurry Organic Chemistry Solutions Manual Scribd Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Mcmurry Organic Chemistry Solutions Manual Scribd
 - Highlighting and Note-Taking Mcmurry Organic Chemistry Solutions Manual Scribd
 - Interactive Elements Mcmurry Organic Chemistry Solutions Manual Scribd

- 8. Staying Engaged with Mcmurry Organic Chemistry Solutions Manual Scribd
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Mcmurry Organic Chemistry Solutions Manual Scribd
- 9. Balancing eBooks and Physical Books Mcmurry Organic Chemistry Solutions Manual Scribd
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Mcmurry Organic Chemistry Solutions Manual Scribd
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Mcmurry Organic Chemistry Solutions Manual Scribd
 - Setting Reading Goals Mcmurry Organic Chemistry Solutions Manual Scribd
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Mcmurry Organic Chemistry Solutions Manual Scribd
 - Fact-Checking eBook Content of Mcmurry Organic Chemistry Solutions Manual Scribd
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
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