

AQA Chemistry

GCSE Exam-style mark scheme

C4 Chemical calculations

100 marks (20%)

1 hour

Question	Answer	Mark	Mark	Mark
10.1.1	1.4 g		1	100
10.1.2	1.48 g	Amount of magnesium carbonate $1.48 \text{ g} \div 142.17 \text{ g mol}^{-1}$ $= 0.01041 \text{ mol}$ $\times 1000 = 10.41 \text{ mmol}$ <u>10</u> mol as amount of Ca^{2+} ions Bromine ions should be produced by 2 moles	10	100
10.1.3	any 1000 factor - weighing error - not all magnesium reacted - not all magnesium oxide used in the reaction		10	100
10.2	all calculations must be shown calculate all ions in the same compound as well as the overall formula mass of oxygen atoms in water (oxygen added separately)	Mass of precipitate of both moles $10.41 \text{ mmol} \times 142.17 \text{ g mol}^{-1}$ $= 1.48 \text{ g}$ Mass of precipitate of water $10.41 \text{ mmol} \times 18.015 \text{ g mol}^{-1}$ $= 0.187 \text{ g}$ Total mass of precipitate $= 1.48 \text{ g} + 0.187 \text{ g}$ $= 1.667 \text{ g}$	10	100
10.3.1	all calculations as above instructions are correct		10	100
10.3.2	Calculation of 1.48 $1.48 \text{ g} \div 142.17 \text{ g mol}^{-1}$ $= 0.01041 \text{ mol}$ $\times 1000 = 10.41 \text{ mmol}$	Mass of precipitate of both moles $10.41 \text{ mmol} \times 142.17 \text{ g mol}^{-1}$ $= 1.48 \text{ g}$ Mass of precipitate of water $10.41 \text{ mmol} \times 18.015 \text{ g mol}^{-1}$ $= 0.187 \text{ g}$ Total mass of precipitate $= 1.48 \text{ g} + 0.187 \text{ g}$ $= 1.667 \text{ g}$	10	100

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