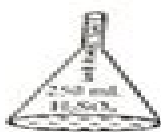


# অধ্যায় ১: ল্যাবরেটরির নিরাপদ ব্যবহার

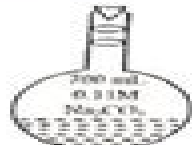


পটাসিয়াম কপ্টার, এখানে মোট পটাসিয়াম, আরও কপ্টারের শীর্ষস্থানীয় কপ্টারের পটাসিয়াম এবং মনোমের টেক্টর পটাসিয়াম  
ন্যূনতম উত্তর অধ্যায়ভিত্তিক মেট্রিক্স হয়েছে। এগুলোর অনুশীলন করলে তুমি এ অধ্যায়টি থেকে সেরাভাবে বুঝতে পারবে।

প্রশ্ন ১



পাত্র-A



পাত্র-B

সি. কে. ১০/১০/১০

- মোলের প্রকণ কী?
- "কেন্দ্র পটাসিয়াম লবণের প্রকণতা, 'অপমাত্রা বৃদ্ধির সাথে সাথে এক পর্যায়ে বৃদ্ধি না পেয়ে হ্রাস পায়'— ব্যাখ্যা করো।
- A-পাত্রের বিন্যাসন এলিটের ভর কত নির্ণয় করো।
- প্রকণ দুটি মিশ্রিত করলে, মিশ্রণটি কী পরিমাণ অক্সিজেন দ্বারা পরিণত হবে— বিশ্লেষণ করো।

১ নং প্রশ্নের উত্তর

1000 mL প্রকণে কোন প্রকণের 1 mol প্রকণিত থাকলে উক্ত প্রকণকে মোলার প্রকণ বলে।

কেন্দ্র পটাসিয়াম লবণের প্রকণতা অপমাত্রা বৃদ্ধির সাথে সাথে এক পর্যায়ে বৃদ্ধি না পেয়ে হ্রাস পায়। এর কারণ, আপমাত্রা বৃদ্ধির সাথে সাথে প্রকণতা বৃদ্ধি পেতে থাকলেও একটি নির্দিষ্ট আপমাত্রার কেন্দ্র পটাসিয়াম লবণ নিষ্কৃতি হয়। তখন প্রকণতা হ্রাস পেতে থাকে। উদাহরণস্বরূপ  $\text{Na}_2\text{SO}_4 \cdot 10\text{H}_2\text{O}$  এর কথা বলা যায়।  $\text{Na}_2\text{SO}_4 \cdot 10\text{H}_2\text{O}$  এর ক্ষেত্রে পটাসিয়ামে প্রকণিত করার সময় বিয়োজন প্রক্রিয়াটি অপমাত্রা বৃদ্ধি 32.4°C আপমাত্রার উপরে  $\text{Na}_2\text{SO}_4 \cdot 10\text{H}_2\text{O}$  যখন  $\text{Na}_2\text{SO}_4$  এ পরিণত হয় তখন বিয়োজন প্রক্রিয়াটি অপোৎপাদী হয়। এ কারণে 32.4°C আপমাত্রার উপরে আপমাত্রা বৃদ্ধিজনিত কারণে প্রকণতার হ্রাস ঘটে।

২

আমরা জানি,  
 $w = SMV$   
 $= (0.5 \times 98 \times 0.25) \text{ g}$   
 $= 12.25 \text{ g}$

অতএব, প্রকণিত  $\text{H}_2\text{SO}_4$  এর ভর  
 $= 12.25 \text{ g}$

উদাহরণের A-পাত্রের অবস্থিতি,  
 $\text{H}_2\text{SO}_4$  এর আণবিক ভর,  
 $M = 1 \times 2 + 32 + 16 \times 4$   
 $= 2 + 32 + 64$   
 $= 98 \text{ g/mol}$

প্রকণের ঘনমাত্রা,  
 $S = 0.5 \text{ M}$   
প্রকণের আয়তন,  
 $V = 250 \text{ mL}$   
 $= \frac{250}{1000} = 0.25 \text{ L}$

প্রকণিত  $\text{H}_2\text{SO}_4$  এর ভর,  $w = ?$

প্রকণ-A:  $\text{H}_2\text{SO}_4$  এর মোল সংখ্যা  $= \left(0.5 \times \frac{250}{1000}\right) \text{ mol}$   
 $= 0.125 \text{ mol}$

প্রকণ-B:  $\text{Na}_2\text{CO}_3$  এর মোল সংখ্যা  $= \left(0.11 \times \frac{250}{1000}\right) \text{ mol}$   
 $= 0.077 \text{ mol}$

সম্প্রতিক বিক্রিয়া:  $\text{H}_2\text{SO}_4 + \text{Na}_2\text{CO}_3 \rightarrow \text{Na}_2\text{SO}_4 + \text{H}_2\text{O} + \text{CO}_2$

সমীকরণ হতে পাই, 1 mol  $\text{H}_2\text{SO}_4 = 1 \text{ mol Na}_2\text{CO}_3$

1 mol  $\text{Na}_2\text{CO}_3$  বিক্রিয়া করে  $\text{H}_2\text{SO}_4$  এর 1 mol এর সাথে

$0.077 \text{ mol} = 0.077 \text{ mol}$   
 $= 0.077 \text{ mol}$

অবশিষ্ট  $\text{H}_2\text{SO}_4$  এর পরিমাণ  $= (0.125 - 0.077) \text{ mol}$   
 $= 0.048 \text{ mol}$

পাত্র-A এর প্রকণের সাথে পাত্র-B এর প্রকণের বিক্রিয়ার ফলে 0.048 mol  $\text{H}_2\text{SO}_4$  অবশিষ্ট থাকায় মিশ্রিত প্রকণের প্রকৃতি হবে অক্সিজেন।

প্রশ্ন ২ একদল ছাত্র ল্যাবরেটরিতে X-পদ্ধতিতে 10 mg এর প্রকণ প্রকণ এবং Y-পদ্ধতিতে 0.5 g পরিমাণ 2%  $\frac{w}{v}$  HCl এর প্রকণ প্রকণ করে।

সি. কে. ১০/১০/১০

- নোট কাকে বলে?
- $\text{Cl}_2$  অপোনার— ব্যাখ্যা করো।
- উদাহরণের এলিটটির মোলার ঘনমাত্রা নির্ণয় করো।
- উদাহরণের কোন পদ্ধতি পরিবেশবান্ধব? বিশ্লেষণ করো।

২ নং প্রশ্নের উত্তর

নিউক্লিয়াসের চারদিকে যে নির্দিষ্ট ত্রিমাত্রিক স্থানে কোন ইলেক্ট্রন পাওয়ার সম্ভাব্যতা শূন্য তাকে নোড বলে।

সমযোজী বন্ধনে আবদ্ধ উভয় পরমাণুর তড়িৎ ঋণাত্মকতা এক হবে অর্থাৎ তড়িৎ ঋণাত্মকতার পার্থক্য শূন্য হবে সমযোজী বন্ধন শূন্য ও অপোনার হয়। আর তড়িৎ ঋণাত্মকতার পার্থক্য 0.5 এর কম হলে অপোনার হয়। যেহেতু Cl এর তড়িৎ ঋণাত্মকতা 3.0 ও C অণুতে তড়িৎ ঋণাত্মকতার পার্থক্য  $(3 - 3) = 0$ , তাই  $\text{Cl}_2$  বিশুদ্ধ সমযোজী ও অপোনার যৌগ। উল্লেখ্য যৌগিক অণুর ক্ষেত্রে এহুণ ঘটে থাকে।

উদাহরণে বর্ণিত এলিটটি হলো 2%  $\left(\frac{w}{v}\right) \text{ HCl}$

2%  $\left(\frac{w}{v}\right) \text{ HCl}$  এর অর্থ হলো—

100 mL প্রকণে HCl এর ভর 2g

∴ 1000 mL " " " "  $\left(\frac{2 \times 1000}{100}\right) \text{ g}$   
 $= 20 \text{ g}$

# Hsc Chemistry 1st Paper

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