



# Hp Officejet J5780 Troubleshooting

**JG Myers**

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web 1 in the given mixture the sum of all the mole fractions is equal to one  $x_a + x_b + \dots = 1$  the multiplication of the mole fraction

by 100 gives the mole percentage 3 mole fraction is a unitless and dimensionless expression advantages of molar fraction  
mole fraction is not dependent on the temperature

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web solution total mass of the answer 100 g mass of benzene 30 g therefore mass of carbon tet 100 30 g benzene s molar mass  $C_6H_6$  6 12 6 1 g mol 1 78 g mol 1 hence  $C_6H_6$  s number of moles  $\frac{30}{78}$  mol 0 3846 mol molar mass of carbon tet  $CCl_4$  1 12 4 355 154 g mol 1

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web solution moles of  $CH_3OH$  5 5 32 0 17 mole moles of  $H_2O$  40 18 2 2 moles therefore according to the equation mole fraction of  $CH_3OH$   $\frac{0.17}{0.17 + 0.17}$  mole fraction of  $CH_3OH$  0 073 to solve more examples on mole fraction formula and practice more question please visit byju s com

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web you may work out the mole fractions on your own problem 5 what is the mass of a sample of a 0 449 molal kbr that contains 2 92 kg of water 1 molality moles solute divided by kilograms solute  $\frac{0.449 \text{ mol}}{1 \text{ kg}} \times 2.92 \text{ kg} \times 119.0023 \text{ g/mol}$  156 g kbr moles times molar mass equals grams

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web calculate the mass percent mole fraction molality and molarity of hexane  $C_6H_{14}$  d 0 660 g cm 3 when it is mixed with

heptane c 7 h 12 d 0 684 g cm<sup>3</sup> in a 35 0 55 0 ml ratio assume that the volumes are additive

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web notice that the mole fraction has no units on it and is written as a decimal value do not change it to percent note of caution you could see the term mole percent it is simply the mole fraction multiplied by 100 for example in the problem just below the mole fraction of cinnamic acid is 0 2885 its mole percent would be 28 85

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web since the mole fraction of the  $O_2$  is 0.050 it contributes 0.050 or if you will 5.0% of the total pressure thus total pressure is 0.2105042 atm problem 9 a sample of 1.43 g of helium and an unweighed quantity of  $O_2$  are mixed in a flask at room temperature

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