

**Assignment #2 – Compound Names and Formulas**  
**Multi-valent ions only**

A. Write the correct formula for the following compounds, all of which have been named using the modern Roman Numeral Method. The ionic charge is given after the first element (metallic) in Roman Numerals.

- copper (II) oxide CuO
- mercury (I) oxide Hg<sub>2</sub>O
- gold (III) chloride AuCl<sub>3</sub>
- thallium (III) bromide TlBr<sub>3</sub>
- bismuth (V) oxide Bi<sub>2</sub>O<sub>5</sub>
- terbium (III) oxide Tb<sub>2</sub>O<sub>3</sub>
- uranium (VI) oxide UO<sub>3</sub>
- protactinium (IV) bromide PaBr<sub>4</sub>
- cerium (III) oxide Ce<sub>2</sub>O<sub>3</sub>
- arsenic (V) sulphide As<sub>2</sub>S<sub>5</sub>

- manganese (III) oxide Mn<sub>2</sub>O<sub>3</sub>
- vanadium (II) bromide VBr<sub>2</sub>
- niobium(V) oxide Nb<sub>2</sub>O<sub>5</sub>
- titanium (III) oxide Ti<sub>2</sub>O<sub>3</sub>
- titanium (III) nitride TiN
- iron (II) oxide FeO
- cobalt (II) phosphide Co<sub>3</sub>P<sub>2</sub>
- tin (II) oxide SnO
- thulium (II) bromide TmBr<sub>2</sub>
- copper (II) bromide CuBr<sub>2</sub>

B. Name the compound using the Roman Numeral Method and balance the ionic charges.

- SnCl<sub>4</sub> Tin (IV) Chloride
- BiBr<sub>5</sub> Bismuth (V) Bromide
- PoO<sub>2</sub> Polonium (IV) Oxide
- PbI<sub>2</sub> Lead (II) Iodide
- HgO Mercury (II) Oxide
- HgCl Mercury (I) Chloride
- Au<sub>2</sub>O<sub>3</sub> Gold (III) Oxide
- FeCl<sub>2</sub> Iron (II) Chloride

- PdF<sub>4</sub> Palladium (IV) Fluoride
- Os<sub>2</sub>O<sub>3</sub> Osmium (III) Oxide
- MoBr<sub>2</sub> Molybdenum (II) Bromide
- VCl<sub>5</sub> Vanadium (V) Chloride
- Mn<sub>2</sub>O<sub>3</sub> Manganese (III) Oxide
- CoO Cobalt (II) Oxide
- Np<sub>2</sub>O<sub>3</sub> Neptunium (III) Oxide
- V<sub>2</sub>O<sub>5</sub> Vanadium (V) Oxide

# Formula Of Ionic Compound Lab Answer Key

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