

ক্র.	বর্ত ও উদ্দেশ্য	সমীকরণ
১	C_p ও C_v এর সম্পর্ক— এখানে, R = মোলার গ্যাস ধ্রুবক	$C_p - C_v = R$ এবং $\frac{C_p}{C_v} = \gamma$
২	n মোল গ্যাসের তাপমাত্রা T_1 থেকে T_2 হলে অণুগত পরিবর্তন—	$\Delta U = nC_v(T_2 - T_1) = nC_v\Delta T$
৩	তাপমাত্রা T স্থির রেখে (সমোক্ত প্রক্রিয়া) n মোল গ্যাসের আয়তন V_1 থেকে V_2 হলে নির্দিষ্ট কর্তৃক পুষ্টিত তাপ বা কৃত কাজ—	$Q = W = nRT \ln \frac{V_2}{V_1}$
৪	সমোক্ত প্রক্রিয়ার চাপ P ও আয়তন V এর মধ্যে সম্পর্ক—	$PV = \text{ধ্রুবক বা } P_1V_1 = P_2V_2$
৫	বৃদ্ধিতাপীয় প্রক্রিয়ার n মোল গ্যাসের তাপমাত্রা T_1 থেকে T_2 হলে নির্দিষ্ট কর্তৃক কৃত কাজ—	$W = nC_v(T_2 - T_1) = n \frac{R}{\gamma - 1} (T_2 - T_1)$
৬	বৃদ্ধিতাপীয় প্রক্রিয়ার চাপ P ও আয়তন V এর মধ্যে সম্পর্ক—	$PV^\gamma = \text{ধ্রুবক বা } P_1V_1^\gamma = P_2V_2^\gamma$
৭	বৃদ্ধিতাপীয় প্রক্রিয়ার তাপমাত্রা T ও আয়তন V এর মধ্যে সম্পর্ক—	$TV^{\gamma-1} = \text{ধ্রুবক বা } T_1V_1^{\gamma-1} = T_2V_2^{\gamma-1}$
৮	বৃদ্ধিতাপীয় প্রক্রিয়ার তাপমাত্রা T ও চাপ P এর মধ্যে সম্পর্ক—	$T^\gamma P^{1-\gamma} = \text{ধ্রুবক বা } T_1^\gamma P_1^{1-\gamma} = T_2^\gamma P_2^{1-\gamma}$
৯	যে কোনো ইঞ্জিন W পরিমাণ কাজ সম্পাদনের সময় Q_1 তাপ গ্রহণ এবং Q_2 তাপ হারান করলে এর দক্ষতা—	$\eta = \frac{W}{Q_1} = 1 - \frac{Q_2}{Q_1}$
১০	কোনো ইঞ্জিন বা প্রত্যাবর্তী ইঞ্জিনের উষ্ণতম তাপমাত্রা T_1 এবং তাপ গ্রহণকের তাপমাত্রা T_2 হলে এর দক্ষতা—	$\eta = 1 - \frac{T_2}{T_1}$
১১	বিষাক্রমের কৃতি পুঙ্খক বা কার্যসম্পাদন সম্বন্ধ— এখানে, Q_1 = নিম্ন তাপমাত্রার তাপাধার থেকে পুষ্টিত তাপ, Q_2 = উচ্চ তাপমাত্রায় পরিণত তাপ এবং W = বাইরে থেকে নির্দিষ্টের উপর কৃতকর্ম বা কৃতকর্ম পরিমাণ—	$\eta = \frac{W}{Q_1} = \frac{Q_1 - Q_2}{Q_1}$
১২	T স্থির তাপমাত্রায় কোনো নির্দিষ্ট A অবস্থা থেকে B অবস্থায় আসতে যদি ΔQ পরিমাণ তাপ শোষণ করে তবে একত্রটির পরিবর্তন—	$\Delta S = S_B - S_A = \frac{\Delta Q}{T}$
১৩	n ভরের কোনো বস্তুর তাপমাত্রা T_1 থেকে T_2 করা হলে এর একত্রটির পরিবর্তন— এখানে, c = আপেক্ষিক তাপ	$\Delta S = \int_{T_1}^{T_2} \frac{mcdT}{T} = mc \ln \frac{T_2}{T_1}$
১৪	n ভরের কোনো বস্তুর তাপমাত্রা T পরিবর্তন না হয়ে অবস্থার পরিবর্তন হলে একত্রটির পরিবর্তন— এখানে, L = ঐ অবস্থা পরিবর্তনের আপেক্ষিক সূত্রতাপ	$\Delta S = \frac{mL}{T}$

এ অধ্যায়ে প্রকৃষ্ট বিভিন্ন ধ্রুবকের মান

তাপমাত্রার স্কেল	সেলসিয়াস
বৈদিক স্বাভাবিক, n	100
নিম্ন শিঃসংকট, θ_{lower}	0°C
উচ্চ শিঃসংকট, θ_{upper}	100°C
মোলার গ্যাস ধ্রুবক, $R = 8.31 \text{ J}\cdot\text{mol}^{-1}\cdot\text{K}^{-1}$	

তাপমাত্রার স্কেল	এক পরমাণুক গ্যাস	দ্বি- পরমাণুক গ্যাস	বহু পরমাণুক গ্যাস
স্থির তাপে মোলার তাপ ধারণ ক্ষমতা, $C_p =$	$20.77 \text{ J}\cdot\text{mol}^{-1}\cdot\text{K}^{-1}$	$29.085 \text{ J}\cdot\text{mol}^{-1}\cdot\text{K}^{-1}$	$33.26 \text{ J}\cdot\text{mol}^{-1}\cdot\text{K}^{-1}$
স্থির আয়তনে মোলার তাপ ধারণ ক্ষমতা, $C_v =$	$12.46 \text{ J}\cdot\text{mol}^{-1}\cdot\text{K}^{-1}$	$20.775 \text{ J}\cdot\text{mol}^{-1}\cdot\text{K}^{-1}$	$24.95 \text{ J}\cdot\text{mol}^{-1}\cdot\text{K}^{-1}$
C_p ও C_v এর অনুপাত, $\gamma =$	1.67	1.4	1.33
পানির আপেক্ষিক তাপ, $c_p =$	$4200 \text{ J}\cdot\text{kg}^{-1}\cdot\text{K}^{-1}$		
বরফের আপেক্ষিক তাপ, $c_f =$	$2100 \text{ J}\cdot\text{kg}^{-1}\cdot\text{K}^{-1}$		
বরফ গলনের আপেক্ষিক সূত্রতাপ, $L_f =$	$336000 \text{ J}\cdot\text{kg}^{-1}$		
পানির বাষ্পীভবনের আপেক্ষিক সূত্রতাপ, $L_v =$	$2260000 \text{ J}\cdot\text{kg}^{-1}$		

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