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Peipei Pang



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$\text{HNO}_3 + \text{Mg}(\text{C}_2\text{H}_3\text{O}_2)_2 \rightarrow \text{H}_2\text{C}_2\text{H}_3\text{O}_2 + \text{Mg}(\text{NO}_3)_2$ chemical equation - Aug 11 2023

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reaction where one mole of solid magnesium Mg and two moles of aqueous nitric acid HNO_3 react to form one mole of aqueous magnesium nitrate $\text{Mg}(\text{NO}_3)_2$ and one mole of dihydrogen H_2 gas show chemical

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web mar 4 2018 $\text{Mg} + 2\text{HNO}_3 \rightarrow \text{Mg}(\text{NO}_3)_2 + \text{H}_2$ on the left side there are Mg 1 magnesium in elemental form H_2 there are 2 molecules of nitric acid each acid has one hydrogen so two molecules of nitric acid is 2 hydrogens N_2 there are 2 molecules of nitric acid each acid has one nitrogen so two molecules of nitric acid is 2 nitrogens O_6

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you balance magnesium and nitric acid you ll need to count all of atoms on both directions of the chemical equation to balance $\text{mg} + \text{hno}_3 \rightarrow \text{mg}(\text{no}_3)_2 + \text{no}_2 + \text{h}_2\text{o}$

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