A 645-gram beaker of water loses 12.1 kj of water. If the water temperature of water started out at 49.5 °C, at what temperature did it end up?

of AT = g/0m-0 AT = -12,000 }9645 g -4,784 /g *C) = -4.48 *C

By Toron + AT + Toron - Thorn = -4.48 °C + 49.5 °C =

45.0 %

8, A 75.0 g piece of glass absorbs 6,350 joules of heat, if its initial temperature is 21.6 °C, what is its final temperature? Its specific heat capacity is 0.840 J/g °C.

48 AT = 0.000-0 AT = 6.350 pt75.0 g - 0.840 p/g CT = 101 °C

86 Tons = 85 + Tons Tons = 100 °C + 27.6 °C =

MALE FIELD

 What is the specific heat capacity of mercury if 25.0 grams of it is heated from 25.0 °C to 155.0 °C while absorbing 455 joules of heat?

#F = F\$5.4 fC + 25.0 fC = 1.80.0 fC

et c = g/0mdf) c = 455 M25,0 g : 130,0 °C) =

0.140 prop/cs

 How much heat is lost when 114 grams of lead cools from 155 °C to 22.2 °C? The specific heat capacity of lead is 0.129 j/g °C.

at at = 22.2 °C = 155 °C = 133 °C

66 g = 1714 gs(0.129 //g *CV-133 *C) =

5.960 For 3.96 kg

High School Chemistry Thermochemistry Questions

Katrin Zwirglmaier

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