

### Problem 1.7

A tank of compressed nitrogen for industrial process use is a cylinder with 6 in. diameter and 4.25 ft length. The gas pressure is 204 atmospheres (gage). Calculate the mass of nitrogen in the tank.

Given: Data on nitrogen tank

Find: Mass of nitrogen

#### Solution

The given or available data is:

$$D = 6\text{-in}$$

$$L = 4.25\text{-ft}$$

$$p = 204\text{-atm}$$

$$T = (59 + 460)\text{-R}$$

$$T = 519\text{ R}$$

$$R_{N_2} = 55.16 \frac{\text{ft}\cdot\text{lb}_f}{\text{lb}\cdot\text{R}} \quad (\text{Table A.6})$$

The governing equation is the ideal gas equation

$$p = \rho \cdot R_{N_2} \cdot T$$

and

$$\rho = \frac{M}{V}$$

where  $V$  is the tank volume

$$V = \frac{\pi}{4} \cdot D^2 \cdot L$$

$$V = \frac{\pi}{4} \times \left( \frac{6}{12} \text{-ft} \right)^2 \times 4.25\text{-ft} = 0.834\text{ ft}^3$$

Hence

$$M = V \cdot \rho = \frac{p \cdot V}{R_{N_2} \cdot T}$$

$$M = 204 \times 14.7 \frac{\text{lb}_f}{\text{in}^2} \times \frac{144\text{-in}^2}{\text{ft}^2} \times 0.834\text{-ft}^3 \times \frac{1}{55.16} \frac{\text{lb}\cdot\text{R}}{\text{ft}\cdot\text{lb}_f} \times \frac{1}{519} \frac{1}{\text{R}} \times 32.2 \frac{\text{lb}\cdot\text{ft}}{\text{s}^2\cdot\text{lb}_f}$$

$$M = 12.6\text{ lb}$$

$$M = 0.391\text{ slug}$$

# **Fox Fluid Mechanics Solution Manual 6th**

**Yijin Wang**



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