Abbreviations.	1	= alternative and acceptable answers for the same marking point
annotations and	10000	= separates marking points
conventions	NOT	= answers which are not worthy of credit
used in the	()	= words which are not essential to gain credit
	200	 (underlining) key words which <u>must</u> be used to gain credit
Mark Scheme	ecf.	= error carried forward
	AW	= alternative wording
	ora	= or reverse argument

Question	Expected answers	Marks
1 (a)	8 1, 2 or 3 p orbitals are OK	[2]
(b)	d orbital 2 ✓ p sub-shell 6 ✓ 3rd shell 18 ✓	[3]
	energy $2s$ $\uparrow \downarrow$ $2p$ $\uparrow \downarrow \uparrow$ \uparrow \uparrow \downarrow \uparrow \downarrow \uparrow \uparrow \downarrow \uparrow \uparrow \downarrow \uparrow \downarrow \uparrow \downarrow \uparrow \uparrow \downarrow \uparrow \uparrow \downarrow \uparrow \uparrow \downarrow \uparrow \uparrow \downarrow \downarrow \uparrow \downarrow \uparrow \downarrow \uparrow \downarrow \downarrow \uparrow \downarrow \uparrow \downarrow \uparrow \downarrow \downarrow \downarrow \downarrow \uparrow \downarrow	
	2s and 2p labels ✓ Ignore any superscripted numbers. 8 electrons in correct levels with arrows correctly shown ✓	[2]
(d) (Energy change when each atom in 1 mole ✓ of gaseous atoms ✓ loses an electron ✓ (to form 1 mole of gaseous 1+ ions). 	[3]
(O²⁺(g) → O³⁺(g) + e⁻ ✓ ✓ 1 mark for correct species; 1 mark for state symbols. No charge required on electron. 	[2]
	Ignore (g) on e	121

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