

Name: \_\_\_\_\_

Date: \_\_\_\_\_

## Student Exploration: Equilibrium and Concentration

**Vocabulary:** chemical equilibrium, concentration, equilibrium, equilibrium constant, reaction quotient, reversible reaction

**Prior Knowledge Questions** (Do these BEFORE using the Gizmo.)

Gary has \$5,000 in his bank account and earns a modest salary. Every month he pays for rent, food, utilities, and entertainment.

- How will Gary's account change if he saves more than he spends? He will have more money in his account.
- How will Gary's account change if he spends more than he saves? He will have less money in his account.
- What happens if Gary spends exactly as much as he saves? He will have no money in his account.

### Gizmo Warm-up

If Gary spends exactly as much as he earns, his savings will be in **equilibrium**. Equilibrium occurs when two opposing processes occur at the same rate, leading to no net change. In the *Equilibrium and Concentration* Gizmo, you will investigate how equilibrium can occur in chemical reactions.

To begin, check that **Reaction 1** is selected. Set **Moles  $\text{NO}_2$**  to 8 and **Moles  $\text{N}_2\text{O}_4$**  to 0.



- Click **Play** (▶) and observe the colliding molecules. What do you notice? The molecules are colliding very slowly.
- In the Gizmo, a blue flash appears every time two reactants combine to form a product. A red flash appears every time a product dissociates into reactants.
- Click **Reset** (↺), and set **Moles  $\text{NO}_2$**  to 0 and **Moles  $\text{N}_2\text{O}_4$**  to 8. Click **Play**.
  - What do you notice now? The amount of liquid has risen.
- When a reaction can proceed in either direction, it is a **reversible reaction**. Based on what you have observed, is the synthesis of  $\text{NO}_2$  into  $\text{N}_2\text{O}_4$  a reversible reaction? Explain.
  - Yes it is because the reactions go either way.

# Explorelearning Equilibrium And Concentration Gizmo Answer Key

**JS Bruner**



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