

Activity A: Alpha decay

Get the Gizmo ready:

- Click Reset (⌂).
- Check that Alpha decay and Uranium are selected.



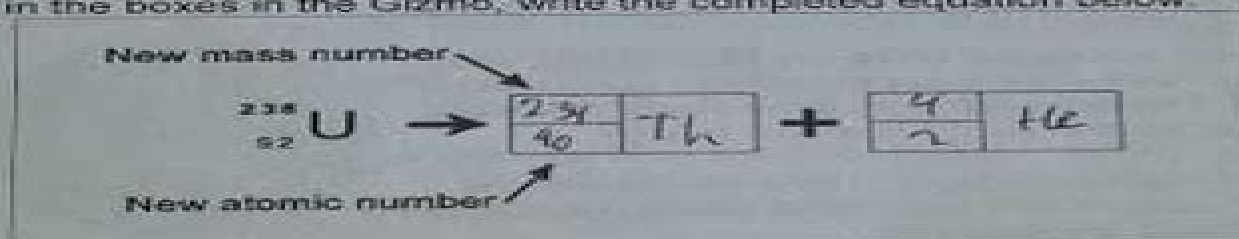
Question: How does alpha decay change the nucleus of a radioactive atom?

Predict: As you observed in the warm-up activity, an alpha particle consists of two protons and two neutrons. How will the emission of an alpha particle affect the following?

- A. The atomic number of the atom: more protons -2
- B. The mass number of the atom: more protons and neutrons -4

Calculate: Turn on **Write equation**. What you see is an equation that shows the original uranium atom on the left. The boxes on the right represent the **daughter product**—the atom produced by radioactive decay—and the emitted alpha particle.

- A. In the top-left box, write the mass number of the daughter product and press "Enter" on your keyboard. What is this number? 234
- B. In the bottom-left box, write the atomic number of the daughter product and press "Enter." What is this number? 90
- C. In the next set of boxes, enter the mass number and atomic number of the alpha particle, which has the same composition as the nucleus of a helium (He) atom. After filling in the boxes in the Gizmo, write the completed equation below:



- D. According to your equation, what version/isotope remains after the alpha decay of uranium-238? (Note: You can look up element symbols on the periodic table.)

The rule -4 = atomic # -2 = mass #

Check: Turn on **Show equation**, and click **Play**. The equation will appear at the end of the animation. Was your prediction correct? yes If not, modify your equation above.

Activity A continued on next page)

Explore Learning Creating An Energy Conversions Answer Key

Daniel F McAuley



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Table of Contents Explore Learning Creating An Energy Conversions Answer Key

1. Understanding the eBook Explore Learning Creating An Energy Conversions Answer Key
 - The Rise of Digital Reading Explore Learning Creating An Energy Conversions Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Explore Learning Creating An Energy Conversions Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Explore Learning Creating An Energy Conversions Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Explore Learning Creating An Energy Conversions Answer Key
 - Personalized Recommendations
 - Explore Learning Creating An Energy Conversions Answer Key User Reviews and Ratings
 - Explore Learning Creating An Energy Conversions Answer Key and Bestseller Lists
5. Accessing Explore Learning Creating An Energy Conversions Answer Key Free and Paid eBooks
 - Explore Learning Creating An Energy Conversions Answer Key Public Domain eBooks
 - Explore Learning Creating An Energy Conversions Answer Key eBook Subscription Services
 - Explore Learning Creating An Energy Conversions Answer Key Budget-Friendly Options

6. Navigating Explore Learning Creating An Energy Conversions Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Explore Learning Creating An Energy Conversions Answer Key Compatibility with Devices
 - Explore Learning Creating An Energy Conversions Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Explore Learning Creating An Energy Conversions Answer Key
 - Highlighting and Note-Taking Explore Learning Creating An Energy Conversions Answer Key
 - Interactive Elements Explore Learning Creating An Energy Conversions Answer Key
8. Staying Engaged with Explore Learning Creating An Energy Conversions Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Explore Learning Creating An Energy Conversions Answer Key
9. Balancing eBooks and Physical Books Explore Learning Creating An Energy Conversions Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Explore Learning Creating An Energy Conversions Answer Key
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Explore Learning Creating An Energy Conversions Answer Key
 - Setting Reading Goals Explore Learning Creating An Energy Conversions Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Explore Learning Creating An Energy Conversions Answer Key
 - Fact-Checking eBook Content of Explore Learning Creating An Energy Conversions Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

-
- Interactive and Gamified eBooks

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