

## Experiment 10

### Oxidation-Reduction (Redox) Reactions

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#### Pre-Lab Assignment

Before coming to lab:

- Read the lab thoroughly.
- Answer the pre-lab questions that appear at the end of this lab exercise.

#### Purpose

A set of oxidation-reduction (redox) reactions will be performed to determine the relative strengths as reducing and oxidizing agents of metals and other ions. An activation series will be devised and used to predict whether a reaction will occur spontaneously or not.

#### Background

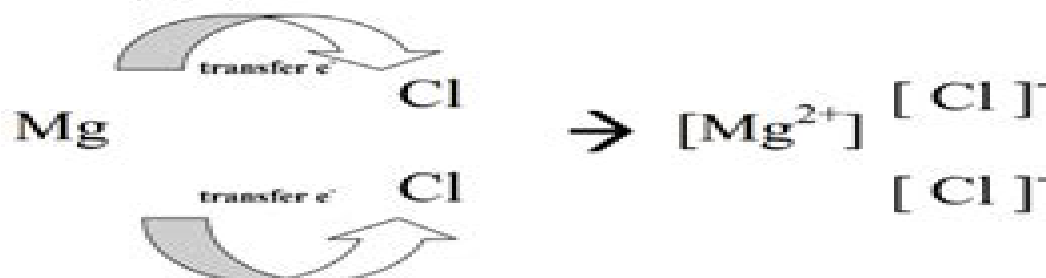
Oxidation-reduction (redox) reactions are a classification of chemical changes that involve the transfer of electrons. An example of a redox reaction is shown in Eqn. 1, when magnesium metal reacts with chlorine gas.



At first glance, the electron exchange is not apparent as electrons are not included in the overall balanced equation. However, electrons are shown when the reaction is split into halves: an oxidized half-reaction and a reduced half-reaction, as in Eqn. 2 and Eqn. 3 respectively.



Notice that in the half-reactions, the electrons are required to balance the charges. Since Mg is becoming more positive, one magnesium atom loses two electrons to become a cation and is said to be **oxidized**. Chlorine is becoming more negative, so one molecule of  $\text{Cl}_2(\text{g})$  gains two electrons to become two chloride anions and is said to be **reduced**. Both oxidation and reduction happen simultaneously in a single redox reaction. Overall, Mg is transferring two electrons to chlorine (Fig. 1).



**Fig. 1:** Redox reaction between Mg(s) and Cl<sub>2</sub>(g) to produce MgCl<sub>2</sub>(aq)

# Experiment 46 Oxidation Reduction Reactions Answers

**JA Banks**



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## **Table of Contents Experiment 46 Oxidation Reduction Reactions Answers**

1. Understanding the eBook Experiment 46 Oxidation Reduction Reactions Answers
  - The Rise of Digital Reading Experiment 46 Oxidation Reduction Reactions Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Experiment 46 Oxidation Reduction Reactions Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Experiment 46 Oxidation Reduction Reactions Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Experiment 46 Oxidation Reduction Reactions Answers
  - Personalized Recommendations
  - Experiment 46 Oxidation Reduction Reactions Answers User Reviews and Ratings
  - Experiment 46 Oxidation Reduction Reactions Answers and Bestseller Lists
5. Accessing Experiment 46 Oxidation Reduction Reactions Answers Free and Paid eBooks

- Experiment 46 Oxidation Reduction Reactions Answers Public Domain eBooks
- Experiment 46 Oxidation Reduction Reactions Answers eBook Subscription Services
- Experiment 46 Oxidation Reduction Reactions Answers Budget-Friendly Options
- 6. Navigating Experiment 46 Oxidation Reduction Reactions Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Experiment 46 Oxidation Reduction Reactions Answers Compatibility with Devices
  - Experiment 46 Oxidation Reduction Reactions Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Experiment 46 Oxidation Reduction Reactions Answers
  - Highlighting and Note-Taking Experiment 46 Oxidation Reduction Reactions Answers
  - Interactive Elements Experiment 46 Oxidation Reduction Reactions Answers
- 8. Staying Engaged with Experiment 46 Oxidation Reduction Reactions Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Experiment 46 Oxidation Reduction Reactions Answers
- 9. Balancing eBooks and Physical Books Experiment 46 Oxidation Reduction Reactions Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Experiment 46 Oxidation Reduction Reactions Answers
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Experiment 46 Oxidation Reduction Reactions Answers
  - Setting Reading Goals Experiment 46 Oxidation Reduction Reactions Answers
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Experiment 46 Oxidation Reduction Reactions Answers
  - Fact-Checking eBook Content of Experiment 46 Oxidation Reduction Reactions Answers
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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