

Enter the % transmission values for each equilibrium mixture obtained in the lab. (Enter measured transmittance to the nearest 0.1%)

Table 1: Experimental Data

Equilibrium Mixture	Volume of $\text{Fe}(\text{NO}_3)_3 \cdot \text{HNO}_3$ Solution Added, mL	%T
1	<input type="text" value="1.0"/>	<input type="text" value="82.3"/>
2	<input type="text" value="2.0"/>	<input type="text" value="68.6"/>
3	<input type="text" value="3.0"/>	<input type="text" value="58.5"/>
4	<input type="text" value="4.0"/>	<input type="text" value="53.7"/>
5	<input type="text" value="5.0"/>	<input type="text" value="49.6"/>
6	<input type="text" value="6.0"/>	<input type="text" value="45.8"/>

# Experiment 3an Equilibrium Constant Answers

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*Nuclear Medicine* Nathan A. Solomon,1977      Equilibrium Calculations John T. Donoghue,1971      **Equilibrium** ,1991

**Chemical Equilibrium** William Guenther,1975-12 The present work is designed to provide a practical introduction to aqueous equilibrium phenomena for both students and research workers in chemistry biochemistry geochemistry and interdisciplinary environmental fields The pedagogical strategy I have adopted makes heavy use of detailed examples of problem solving from real cases arising both in laboratory research and in the study of systems occurring in nature The procedure starts with mathematically complete equations that will provide valid solutions of equilibrium problems instead of the traditional approach through approximate concentrations and idealized infinite dilution assumptions There is repeated emphasis on the use of corrected conditional equilibrium constants and on the checking of numerical results by substitution in complete equations and or against graphs of species distributions Graphical methods of calculation and display are used extensively because of their value in clarifying equilibria and in leading one quickly to valid numerical approximations The coverage of solution equilibrium phenomena is not however exhaustively comprehensive Rather I have chosen to offer fundamental and rigorous examinations of homogeneous step equilibria and their interactions with solubility and redox equilibria Many examples are worked out in detail to demonstrate the use of equilibrium calculations and diagrams in various fields of investigation      **Determination of the Equilibrium Constant of the Reaction  $2\text{A1(1)} + \text{A1C13(g)}$**  H. Mitani,1967

BASIC Equilibrium Calculations Cooper Harold Langford,1973      **Determination of an Equilibrium Constant by Spectrophotometry** James M. Postma,Julian Roberts, Jr.,J. Leland Hollenberg,2000-01-15      Principles of Chemical Equilibrium Kelso B. Morris,1965      *Interpreting an Equilibrium Constant* ,2016 Your examination of the equilibrium constant continues Learn exactly what the numerical value for an equilibrium constant tells and doesn't tell you about an equilibrium system      *Determining an Equilibrium Constant Using Spectrophotometry* Norman J. Hudak,1988-01-01

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Commission on equilibrium data,      **The Equilibrium Constant in the System:  $\text{Fe}^{+++} + \text{Ag}$**  Fred Miller Roberts,1941

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