

Directions: Read carefully and solve the following problems.

1. The complete combustion of Octane results to the production of Carbon dioxide and water as shown in the unbalanced chemical equation below.



- Write the balanced chemical reaction equation.
 - How much CO_2 in grams will be produced from burning 30 grams of Octane?
 - What mass of Oxygen is needed to completely consume 100 grams of Octane?
2. In a certain experiment, Hydrochloric acid is used to decompose Calcium carbonate. The chemical reaction equation is shown below.



- After the experiment, the total mass of CaCl_2 produced is 120g. Assuming that the reaction proceeded to completion, how many moles of CaCO_3 were used in the experiment?
 - Based on the amount of CaCl_2 produced in the reaction, how much CO_2 in grams effervesced out of the Calcium carbonate used?
3. Hydrogen fluoride is produced by reacting Calcium fluoride with sulfuric acid. A by-product of the process is Calcium sulfate.

Exercise 11 Stoichiometry

**Azariah Thomas Lincoln, James Henri
Walton**



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