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Thermodynamics

- Which of the following statements is correct for the spontaneous absorption of a gas? [2014]
 - ΔS is negative and therefore, ΔH should be highly positive
 - ΔS is negative and therefore, ΔH should be highly negative
 - ΔS is positive and therefore, ΔH should be negative
 - ΔS is positive and therefore, ΔH should also be highly positive
- For the reaction, $X_2O_4(l) \longrightarrow 2XO_2(g)$, $\Delta U = 2.1 \text{ kcal}$, $\Delta S = 20 \text{ cal K}^{-1}$ at 300 K. Hence, ΔG is [2014]
 - 2.7 kcal
 - 2.7 kcal
 - 9.3 kcal
 - 9.3 kcal
- For a given exothermic reaction, K_p and K_p' are the equilibrium constants at temperatures T_1 and T_2 , respectively. Assuming that heat of reaction is constant in temperature range between T_1 and T_2 , it is readily observed that [2014]
 - $K_p > K_p'$
 - $K_p < K_p'$
 - $K_p = K_p'$
 - $K_p = \frac{1}{K_p'}$
- A reaction having equal energies of activation for forward and reverse reactions has [2013]
 - $\Delta S = 0$
 - $\Delta G = 0$
 - $\Delta H = 0$
 - $\Delta H = \Delta G = \Delta S = 0$
- In which of the following reactions, standard reaction entropy changes (ΔS°) is positive and standard Gibbs energy change (ΔG°) decreases sharply with increasing temperature? [2012]
 - $C(\text{graphite}) + \frac{1}{2}O_2(g) \longrightarrow CO(g)$
 - $CO(g) + \frac{1}{2}O_2(g) \longrightarrow CO_2(g)$
 - $Mg(s) + \frac{1}{2}O_2(g) \longrightarrow MgO(s)$
 - $\frac{1}{2}C(\text{graphite}) + \frac{1}{2}O_2(g) \longrightarrow \frac{1}{2}CO_2(g)$
- The enthalpy of fusion of water is 1.435 kcal/mol. The molar entropy change for the melting of ice at 0°C is [2012]
 - 10.52 cal/mol K
 - 21.04 cal/mol K
 - 5.260 cal/mol K
 - 0.526 cal/mol K
- Standard enthalpy of vaporisation $\Delta_{\text{vap}}H^\circ$ for water at 100°C is 40.66 kJ mol⁻¹. The internal energy of vaporisation of water at 100°C (in kJ mol⁻¹) is (assume water vapour to behave like an ideal gas). [2012]
 - + 37.56
 - 43.76
 - + 43.76
 - + 40.66
- If the enthalpy change for the transition of liquid water to steam is 30 kJ mol⁻¹ at 27°C, the entropy change for the process would be [2011]
 - 1.0 J mol⁻¹K⁻¹
 - 0.1 J mol⁻¹K⁻¹
 - 100 J mol⁻¹K⁻¹
 - 10 J mol⁻¹K⁻¹

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