

GLOBAL EDITION

Meriam's **ENGINEERING MECHANICS:** DYNAMICS

SI VERSION

BASED ON THE
NINTH EDITION

**J. L. MERIAM
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web limiting reagents and percentage yield worksheet 1 consider the reaction $\text{I}_2\text{O}_5 + 5\text{CO} \rightarrow 5\text{CO}_2 + \text{I}_2$ a 80.0 grams of iodine v oxide I_2O_5 reacts with 28.0 grams of carbon monoxide CO determine the mass of iodine I_2 which could be produced 80 g I_2O_5 1 mol I_2O_5 1 mol I_2 xs 1 333.8 g I_2O_5 1 mol I_2O_5 28 g CO 1 mol CO

7.3 limiting reactant and percent yield problems - Jun 13 2023

web jun 2 2020 the percent yield of a reaction is the ratio of the actual yield to the theoretical yield multiplied by 100 to give a percentage text percent yield text actual yield g over text theoretical yield g times 100 label 3 7 3