

Name: _____

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Ch 11 Worksheet 5 - Thermochemical Stoichiometry

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1. Given the equation: $2 \text{H}_2(\text{g}) + \text{O}_2(\text{g}) \rightarrow 2 \text{H}_2\text{O}(\text{l})$ $\Delta \text{H} = -571.7 \text{ kJ}$
- How much energy is produced if 100.0 grams of H_2 are converted to water?
 - How many grams of H_2 would be required to produce 5000.0 kJ?
2. Given the equation: $\text{N}_2(\text{g}) + 3\text{H}_2(\text{g}) \rightarrow 2\text{NH}_3(\text{g})$ $\Delta \text{H}_{\text{rxn}} = -92.22 \text{ kJ}$
- Is the formation of ammonia endothermic or exothermic?
 - What is the enthalpy change associated with the formation of 1.0 mol of NH_3 ?
 - Is the *decomposition* of ammonia endothermic or exothermic?
 - What is the enthalpy change associated with the decomposition of 4.00 mol NH_3 ?
 - How much energy will be released when 12.0 g $\text{H}_2(\text{g})$ react with excess $\text{N}_2(\text{g})$ to form $\text{NH}_3(\text{g})$?
3. Given the equation: $2\text{Ca}(\text{s}) + \text{O}_2(\text{g}) \rightarrow 2\text{CaO}(\text{s})$ $\Delta \text{H} = -635.09 \text{ kJ}$
- How much energy is released during the production of 5.0 g of CaO ?
 - How much energy must be added to CaO to produce 64.0 g of $\text{O}_2(\text{g})$?
4. Given the equation: $4\text{Fe}(\text{s}) + 3 \text{O}_2(\text{g}) \rightarrow 2\text{Fe}_2\text{O}_3(\text{s})$ $\Delta \text{H} = -824.2 \text{ kJ}$
- Assuming that the iron is converted to iron (III) oxide as in the above equation, how much heat can be provided by a hot pack that contains 50.0 g of iron?
 - If the hot pack from part a. gives off heat at the rate of 42.0 kJ /hour, how long will it stay warm?

Answers:

1a. 14,150 kJ produced

1b. 35.3g

2a.

2b. -46kJ

2c.

2d. 184 kJ

2e. 183 kJ

3a. 28 kJ released

3b. 1270 kJ

4a. 185 kJ provided

4b. 4.40 hours

Energy Ws 5 Thermochemical Equations And Stoichiometry

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