

Worksheet – Calorimetry

Calorimetry is the experimental measurement of heat (q) produced in chemical and physical processes. Heat can not be measured directly, but temperature changes can be measured. The factor that links these two is **heat capacity**.

Heat capacity, C , is defined as the heat required to raise the temperature of a substance one degree. It relates a measured change in temperature (ΔT) to heat flow (q) by the equation:

$$q = C \times \Delta T \times \text{mass}$$

This can be calculated on a per gram basis (specific heat capacity) or a per mole basis (molar heat capacity).

It is easiest to understand the molecular basis for heat capacity by using molar heat capacities of different substances. As energy enters a system, it is distributed among all of the degrees of freedom within a molecule (translational, rotational and vibrational modes). The more complex the molecule, the larger its heat capacity, and the smaller the rise in temperature, since the energy is spread out among more modes.

1. 10 gram samples of iron at a temperature of 100°C are placed into three beakers each containing 1.0 mol of three different solvents, octane, trichloromethane and methanol, all at 25.0°C . Their structures and molar masses are shown below.



octane C_8H_{18}
($\text{mw} = 114 \text{ g/mol}$)



trichloromethane CHCl_3
($\text{mw} = 119 \text{ g/mol}$)



methanol CH_3OH
($\text{mw} = 32 \text{ g/mol}$)

- How many degrees of freedom in each compound? $3 \times n$ $n = \# \text{ atoms}$
octane $26 \times 3 = 78$ trichloromethane $3 \times 5 = 15$ methanol $3 \times 6 = 18$
- Which will have the largest molar heat capacity? *most modes - octane*
- This compound have the (largest, smallest) change in temperature.
- Which will have the largest change in temperature? *smallest C = trichloro
methane*
- The temperature of the methanol rises to 28.9°C . The specific heat capacity of iron is $0.449 \text{ J/g}^\circ\text{C}$. What are the **specific and molar heat capacities** of methanol?
 $q_{\text{metal}} = -q_{\text{solvent}}$ $1 \text{ mol} = 32 \text{ g}$

$$\begin{aligned}
 (0.449 \text{ J/g}^\circ\text{C})(10 \text{ g})(28.9 - 100) &= -(C)(32 \text{ g})(28.9 - 25) \\
 C &= \frac{2.56 \text{ J}}{\text{g}^\circ\text{C}} \times \frac{32 \text{ g}}{\text{mol}} = 81.9 \text{ J/mol}^\circ\text{C}
 \end{aligned}$$

Energy Pogil Answers Calorimetry

Daniel F McAuley



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