

Name _____ Date _____ Class _____

CHAPTER 16 STUDY GUIDE FOR CONTENT MASTERY**Energy and Chemical Change****Section 16.1 Energy**

In your notebook, read about the nature of energy.

In the space at the left, write true if the statement is true; if the statement is false, change the italicized word or phrase to make it true.

- true* _____ 1. Energy is the ability to do work or produce heat.
- cannot be* _____ 2. The law of conservation of energy states that energy can be created and destroyed.
- true* _____ 3. Chemical potential energy is energy stored in a substance because of its composition.
- true* _____ 4. Heat is a form of energy that flows from a warmer object to a colder object.
- true* _____ 5. A calorie is the amount of energy required to raise the temperature of one gram of pure water by one degree Celsius.
- false* _____ 6. A calorie is the SI unit of heat and energy.
- true* _____ 7. The specific heat of a substance is the amount of heat required to raise the temperature of one gram of that substance by one degree Celsius.

- true* _____ 8. Kinetic energy is energy of motion.
- both potential and kinetic energy* _____ 9. Chemical potential energy is a chemical reaction *contains only potential energy*.
- 1000 calories* _____ 10. One nutritional Calorie is equal to 100 calories.

- true* _____ 11. One calorie equals 4.184 joules.
- true* _____ 12. When a fuel is burned, some of its chemical potential energy is lost as heat.
- multiply* _____ 13. To convert kilojoules to joules, divide the number of kilojoules by 1000 joules/1 kilojoule.

Answer the following questions. Show all your work.

14. If the temperature of a 500-g sample of liquid water is raised 2.00°C, how much heat is absorbed by the water? The specific heat of liquid water is 4.184 J/g·°C.

$$q = c \times m \times \Delta T$$

$$q = 4.184 \text{ J/g} \cdot ^\circ\text{C} \times 500 \text{ g} \times 2.00^\circ\text{C}$$

$$q = 4.184 \times 10^3 \text{ J} = 4.18 \text{ kJ}$$

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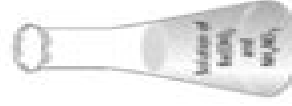
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CHAPTER 16 STUDY GUIDE FOR CONTENT MASTERY**Section 16.2 Heat in Chemical Reactions and Processes**

In your notebook, read about measuring heat and about chemical energy and the universe.

For each item in Column A, write the letter of the matching item in Column B.

- | Column A | Column B |
|--|--------------------------------|
| b. An insulated device used to measure the amount of heat absorbed or released during a chemical or physical process | a. system |
| c. The study of heat changes that accompany chemical reactions and phase changes | b. calorimeter |
| d. The specific part of the universe that contains the reaction or process you wish to study | c. thermodynamics |
| e. The change in enthalpy in a chemical reaction | d. universe |
| f. A system plus its surroundings | e. enthalpy |
| g. The heat content of a system at constant pressure | f. enthalpy change of reaction |
| h. Everything in the universe except the system being studied | g. surroundings |



Use the illustration to answer the following questions.

9. A solution is studying the solution in the flask. What is the system?
The system is the solution of NaOH and H₃PO₄.
10. What are the surroundings?
The surroundings include everything except the solution.
11. What is the universe?
The universe is the solution plus the surroundings.

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Leo J. Malone, Theodore O. Dolter



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