

Empirical and Molecular Formulas Worksheet

Key

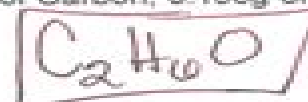
Objectives:

- be able to calculate empirical and molecular formulas

Empirical Formula

- 1) What is the empirical formula of a compound that contains 0.783g of Carbon, 0.196g of Hydrogen and 0.521g of Oxygen?

$$\begin{array}{l} \text{C} \quad \frac{0.783\text{g C}}{12.01\text{g}} = \frac{0.065}{0.033} \approx 1.96 \quad 2 \\ \text{H} \quad \frac{0.196\text{g H}}{1.01\text{g}} = \frac{0.194}{0.033} \approx 5.88 \quad 6 \\ \text{O} \quad \frac{0.521\text{g O}}{16.00\text{g}} = \frac{0.033}{0.033} = 1 \end{array}$$



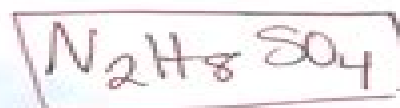
- 2) What is empirical formula of a compound which consists of 89.14% Au and 10.80% of O?

$$\begin{array}{l} \text{Au} \quad \frac{89.14\text{g}}{196.97} = \frac{0.45}{0.45} = 1 \\ \text{O} \quad \frac{10.80\text{g}}{16\text{g}} = \frac{0.675}{0.45} = 1.5 \end{array}$$



- 3) What is empirical formula if compound consists of 21.2%N, 6.1%H, 24.2%S and 48.5%O?

$$\begin{array}{l} \text{N} \quad \frac{21.2\text{g}}{14.01} = 1.5 \quad 2 \\ \text{H} \quad \frac{6.1\text{g}}{1.01\text{g}} = 6.03 \quad 8 \\ \text{S} \quad \frac{24.2\text{g}}{32} = 0.75 \quad 1 \\ \text{O} \quad \frac{48.5\text{g}}{16.00\text{g}} = 3.03 \quad 4 \end{array}$$



Molecular Formula

- 4) Empirical formula of a substance is CH_2O . Molar mass is 180. What is the molecular formula?

$$\text{EFM} = (12 + 2 \times 1 + 16) = 30$$

$$\frac{180}{30} = 6$$



- 5) Sample (3.585g) contains 1.388g of C, 0.345g of H, 1.850g O and its molar mass is 62g. What is molecular formula of this substance?

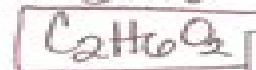
$$\text{C} \quad \frac{1.388}{3.585} \rightarrow 38.7\% \rightarrow \frac{3.87\text{g}}{12.01\text{g}} = \frac{3.22}{3.22} = 1$$

$$\text{H} \quad \frac{0.345}{3.585} \rightarrow 9.62\% \rightarrow \frac{9.62\text{g}}{1.01\text{g}} = \frac{9.52}{3.22} = 3$$

$$\text{O} \quad \frac{1.850}{3.585} \rightarrow 51.6\% \rightarrow \frac{51.6\text{g}}{16\text{g}} = \frac{3.22}{3.22} = 1$$

$$\text{CH}_3\text{O} = \text{EFM} (12 + 3 \times 1 + 16)$$

$$\frac{62}{31.03} = 2$$



1	$\text{C}_2\text{H}_6\text{O}$
2	Au_2O_3
3	NH_4SO_4
4	$\text{C}_2\text{H}_6\text{O}_2$
5	$\text{C}_6\text{H}_{12}\text{O}_6$

Empirical Molecular Formula Answers

Mike Smith

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the name bar.

Empirical Molecular Formula Answers:

Understanding General Chemistry Atef Korchef, 2022-03-07 *Understanding General Chemistry* details the fundamentals of general chemistry through a wide range of topics relating the structure of atoms and molecules to the properties of matter. Written in an easy to understand format with helpful pedagogy to fuel learning, the book features main objectives at the beginning of each chapter, get smart sections, and check your reading section at the end of each chapter. The text is filled with examples and practices that illustrate the concepts at hand. In addition, a summary and extensive MCQs, exercises, and problems with the corresponding answers and explanations are readily available. Additional features include Alerts students to common mistakes and explains in simple ways and clear applications how to avoid these mistakes. Offers answers and comments alongside sample problems enabling students to self evaluate their skill level. Includes powerful methods, easy steps, simple and accurate interpretations, and engaging applications to help students understand complex principles. Provides a bridge to more complex topics such as solid state chemistry, organometallic chemistry, chemistry of main group elements, inorganic chemistry, and physical chemistry. This introductory textbook is ideal for chemistry courses for non science majors as well as health sciences and preparatory engineering students. **CBSE Chemistry Chapterwise Case Study Class 11**

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