Time: 3 hrs.l

MARCH - 2015

Note: (i) All questions are compulsory.

(a) $\Delta G < 0$

Answer to the two sections should be written in the same answer book.

(iii) Figures to the right hand side indicate full marks.

(iv) Write balanced chemical equations and draw neat and labelled diagrams wherever necessary.

(v) Use of logarithmic table is allowed.

(vi) Answer to every new question must be started on a new page.

SECTION - I

Q. 1. Select and write the most appropriate answers from the given alternatives for each sub-question:

[7]

(i) p-type semi-conductors are made by mixing silicon with impurities of _ (a) germanium (b) boron (c) arsenic (d) antimony

(ii) Amongst the following identify the criterion for a process to be at equilibrium :

(c) $\Delta S_{total} = 0$ (d) $\Delta S < 0$ (iii) Colligative property depends only on ___ in a solution.

(a) Number of solute particles. (b) Number of solvent particles.

(c) Nature of solute particles. (d) Nature of solvent particles.

(iv) The charge of how many coulomb is required to deposit 1.0 g of sodium metal (molar mass 23.0 g mol-1) from sodium ions is (c) 193000

(a) 2098 (b) 96500

(b) $\Delta G > 0$

(d) 4196

(v) What is the chemical composition of malachite? (a) CuO-CuCO₃ (b) Cu(OH)₂-CuCO₃ (c) CuO-Cu(OH)₂ (d) Cu₂O-Cu(OH)₂

(vi) The element that does NOT exhibit allotropy is ___

(a) As (d) N (b) Sb (c) Bi

(vii) The integrated rate equation for first order reaction is A $\xrightarrow{\text{products}}$ (a) $K = 2.303 \, \text{t} \, \log_{10} \frac{\left[\mathring{A}\right]_0}{\left[\mathring{A}\right]_t}$ (b) $K = -\frac{1}{t} \, \ln \frac{\left[\mathring{A}\right]_t}{\left[\mathring{A}\right]_0}$

(c) $K = \frac{2.303}{t} \log_{10} \frac{[A]_t}{[A]}$ (d) $K = \frac{1}{t} \ln \frac{[A]_t}{[A]_t}$

Q. 2. Answer any SIX of the following:

[12]

(i) Define the following terms:

(a) Enthalpy of fusion (b) Enthalpy of atomization

Derive van't Hoff general solution equation.

(iii) Explain impurity defect in stainless steel with diagram.

(iv) Derive the relation between half life and rate constant for a first order reaction.

(v) Draw neat and labelled diagram of dry cell.

(vi) Explain the structure of sulphur dioxide.

(vii) What is calcination? Explain it with reactions.

(viii) Arrange the following reducing agents in the order of increasing strength under

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- 1. Understanding the eBook Hsc Chemistry Paper March 2015 Mistakes
 - The Rise of Digital Reading Hsc Chemistry Paper March 2015 Mistakes
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Hsc Chemistry Paper March 2015 Mistakes
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Hsc Chemistry Paper March 2015 Mistakes
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Hsc Chemistry Paper March 2015 Mistakes
 - Personalized Recommendations

- Hsc Chemistry Paper March 2015 Mistakes User Reviews and Ratings
- Hsc Chemistry Paper March 2015 Mistakes and Bestseller Lists
- 5. Accessing Hsc Chemistry Paper March 2015 Mistakes Free and Paid eBooks
 - Hsc Chemistry Paper March 2015 Mistakes Public Domain eBooks
 - Hsc Chemistry Paper March 2015 Mistakes eBook Subscription Services
 - Hsc Chemistry Paper March 2015 Mistakes Budget-Friendly Options
- 6. Navigating Hsc Chemistry Paper March 2015 Mistakes eBook Formats
 - ∘ ePub, PDF, MOBI, and More
 - Hsc Chemistry Paper March 2015 Mistakes Compatibility with Devices
 - Hsc Chemistry Paper March 2015 Mistakes Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Hsc Chemistry Paper March 2015 Mistakes
 - Highlighting and Note-Taking Hsc Chemistry Paper March 2015 Mistakes
 - Interactive Elements Hsc Chemistry Paper March 2015 Mistakes
- 8. Staying Engaged with Hsc Chemistry Paper March 2015 Mistakes
 - o Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - $\circ\,$ Following Authors and Publishers Hsc Chemistry Paper March 2015 Mistakes
- 9. Balancing eBooks and Physical Books Hsc Chemistry Paper March 2015 Mistakes
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Hsc Chemistry Paper March 2015 Mistakes
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Hsc Chemistry Paper March 2015 Mistakes
 - Setting Reading Goals Hsc Chemistry Paper March 2015 Mistakes
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Hsc Chemistry Paper March 2015 Mistakes
 - Fact-Checking eBook Content of Hsc Chemistry Paper March 2015 Mistakes

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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