পরীকারী কপুরা, এখানে বোর্র পরীকা, ক্যাডোট কলেলগর শীর্ষান্ধানীর কলেলগমূহের পরীক্ষা এবং মহেল টেন্টের ব্যাত্তর পূর্ণাল উত্তর অধ্যায়ভিত্তিক দেবছা ময়েছে। এগুলো অনুশীলন করনে কৃষি এ অধ্যায়টি লেকে সেকোনো স্থানশীল জনামূদ্

#### 





ক, মোলার দ্রবণ কীৰু

 "কেলাস পানিযুক্ত লকদের দ্রাক্তার, তাপমাত্রা বুল্পির সাথে সাথে এক পর্যায়ে বৃশ্বি না পেয়ে হ্রাস পায়"— ব্যাখ্যা করে।

ণ. A-পাত্রে বিদর্শান এগিজের ভর কত নির্ণয় করো।

 ছবল দুটি মিশ্রিত করলে, থিপ্রদটি কী পরিমাণ অগ্নীয় বা ফারীয় মবে— বিশ্লেষণ করো।

#### ১ নং প্রয়ের উত্তর

1000 ml. চবণে কোন স্তবের 1 mol স্তবীভূত থাকলে উদ্ভ স্তবগকে
মোলার স্তবণ কলে।

ত্রনাস পানিত্র গবণের প্রাব্যান্তা তাপমাত্রা বৃশ্বির সাথে সাথে এক পর্যান্ত বৃশ্বির না পেরা ছাস পার। এর কারণ, তাপমাত্রা বৃশ্বির সাথে সাথে রাক বারণ, তাপমাত্রা বৃশ্বির সাথে সাথে প্রকাশ রাক্তি বিশ্বিন্ত তাপমাত্রার কেলাস পানি কুর গবণ নির্দ্ধিন হর। তথন প্রাব্যান্তা স্ত্রাস পেতে থাকে। উদ্দেশ্যমন্ত্রপ Na<sub>2</sub>SO<sub>4</sub>, 10H<sub>2</sub>O এর কথা করা যায়। Na<sub>2</sub>SO<sub>4</sub>, 10H<sub>2</sub>O এর পেত্রে পানিতে প্রবীত্ত হওয়ার সময় বিয়োজন প্রক্রিয়ান্তি ভাপহারী বিক্রিয়া কিন্তু 32,4°C তাপমাত্রার উপরে Na<sub>2</sub>SO<sub>4</sub>,10H<sub>2</sub>O যখন Na<sub>2</sub>SO<sub>4</sub> পরিণত হর তথন বিয়োজন প্রক্রিয়াটি ভাপোৎপাদী হয়। এ কারণে 32,4°C তাপমাত্রার উপরে তাপমাত্রা কৃষ্ণিজনিত কারণে প্রাব্যান্তর স্থান হল হল হল হলে।

## कामदा कानि

w = SMV = (0.5 × 98 × 0.25) g = 12.25 g অভন্তন্ত্র, দ্রবীভূত H<sub>2</sub>SO<sub>4</sub> এর ভর

উদ্দীপকের A-পাত্রে অবস্থিত,

H<sub>2</sub>SO<sub>4</sub> এর আণবিক ভর,

M = 1 × 2 + 32 + 16 × 4

= 2 + 32 + 64

= 98 g/mol

রবণের ঘনমারা,

S = 0.5 M

রবণের আয়তন,

V = 250 mL

= 250

1000 = 0.25L

ফবীভূত H<sub>2</sub>SO<sub>4</sub> এর ভর, w = 1

ভাল-A: H<sub>2</sub>SO<sub>4</sub> গান্ত দোলা সংখ্যা —  $\left(0.5 \times \frac{250}{1000}\right)$  med — 0.125 mod — 0.125 mod — 0.125 mod — 0.077 mod — 0.077 mod — 0.077 mod সংঘটিত বিভিন্না: H<sub>2</sub>SO<sub>4</sub> + Na<sub>2</sub>CO<sub>3</sub> → Na<sub>2</sub>SO<sub>4</sub> + H<sub>2</sub>O + CO<sub>3</sub> স্থীকরণ হতে পাই, I mod H<sub>2</sub>SO<sub>4</sub> = I mod Na<sub>3</sub>CO<sub>3</sub> I mod Na<sub>2</sub>CO<sub>3</sub> বিভিন্না করে H<sub>2</sub>SO<sub>4</sub> এই I mod এই সংস্থা —  $\left(0.077\right)$  mod — —  $\left(0.077$ 

19.05

ett.

পাত্র–∧ এর স্তবণের সাথে পাত্র–৪ এর প্রবণের বিভিন্নার ছলে ১৯ mol II₂SO, অবশিউ থাকায় মিশ্রিত স্তবণের প্রকৃতি ছবে অভীত।

ভর্মা > ব্যক্তন ছাত্র লাবরেটরিতে X-পশ্বতিতে 10 mg বর তেওঁ হ পরিমাণ এবং Y-পশ্বতিতে 0.5 g পরিমাণ 2% ₩ HCl এর নমূর হ কাজ করে।

ক. লোভ কাকে বলে?

খ. Cl<sub>2</sub> অপোলার- ব্যাখ্যা করে।

ণ, উন্দীপকের এসিভটির মোলার ঘনমাতা নির্ণয় করে।

য়, উদ্দীপকের কোন পশ্বতি পরিবেশবান্ধর? বিশ্লেষণ করে। ২ নং প্রয়ের উত্তর

নিউরিয়াসের চারলিকে যে নির্দিট ত্রিমাত্রিক স্থানে কানে ইজক পাওয়ার সম্ভাব্যতা শূন্য তাকে নোভ বলে।

সমযোজী বন্ধনে আবন্ধ উভয় পরমাণুর তড়িং কণাছকত ক হলে অর্থাৎ তড়িং ঝণাঝকতার পার্থকা শূন্য হলে সমযোজী ক বিশুন্ধ ও অপোলার হয়। আর তড়িং ঝণাঝকতার পার্থকা 0.5 এর ক মর্দে অপোলার হয়। মেহেতু CI এর তড়িং ঝণাঝকতা 3.0 ৫ ট অপুতে তড়িং ঝণাঝকতার পার্থকা (3 – 3)= 0, তাই CI<sub>2</sub> বিশ্ব সমযোজী ও অপোলার যৌগ। উল্লেখ্য মৌলিক অপুর ক্ষেত্রে এমুশ হা থাকে।

ত্রী উদ্দীপকে বর্ণিত এগিচটি হলো 2%  $\left(\frac{W}{V}\right)$  HCI ।

2%  $\left(\frac{W}{V}\right)$  HCI এর অর্থ ফলো- 100 mL দ্রবংশ HCI এর ভর  $2\epsilon$ 

.. 1000 mL " " (2 × 1000) 100 g

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## **Table of Contents Hsc Chemistry 1st Paper Question 2014 Out**

- 1. Understanding the eBook Hsc Chemistry 1st Paper Question 2014 Out
  - $\circ\,$  The Rise of Digital Reading Hsc Chemistry 1st Paper Question 2014 Out
  - Advantages of eBooks Over Traditional Books
- 2. Identifying Hsc Chemistry 1st Paper Question 2014 Out
  - Exploring Different Genres
  - o Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Hsc Chemistry 1st Paper Question 2014 Out
  - User-Friendly Interface
- 4. Exploring eBook Recommendations from Hsc Chemistry 1st Paper Question 2014 Out
  - Personalized Recommendations
  - Hsc Chemistry 1st Paper Question 2014 Out User Reviews and Ratings

- Hsc Chemistry 1st Paper Question 2014 Out and Bestseller Lists
- 5. Accessing Hsc Chemistry 1st Paper Question 2014 Out Free and Paid eBooks
  - Hsc Chemistry 1st Paper Question 2014 Out Public Domain eBooks
  - Hsc Chemistry 1st Paper Question 2014 Out eBook Subscription Services
  - Hsc Chemistry 1st Paper Question 2014 Out Budget-Friendly Options
- 6. Navigating Hsc Chemistry 1st Paper Question 2014 Out eBook Formats
  - o ePub, PDF, MOBI, and More
  - Hsc Chemistry 1st Paper Question 2014 Out Compatibility with Devices
  - Hsc Chemistry 1st Paper Question 2014 Out Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Hsc Chemistry 1st Paper Question 2014 Out
  - Highlighting and Note-Taking Hsc Chemistry 1st Paper Question 2014 Out
  - Interactive Elements Hsc Chemistry 1st Paper Question 2014 Out
- 8. Staying Engaged with Hsc Chemistry 1st Paper Question 2014 Out
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Hsc Chemistry 1st Paper Question 2014 Out
- 9. Balancing eBooks and Physical Books Hsc Chemistry 1st Paper Question 2014 Out
  - Benefits of a Digital Library
  - o Creating a Diverse Reading Collection Hsc Chemistry 1st Paper Question 2014 Out
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Hsc Chemistry 1st Paper Question 2014 Out
  - Setting Reading Goals Hsc Chemistry 1st Paper Question 2014 Out
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Hsc Chemistry 1st Paper Question 2014 Out
  - Fact-Checking eBook Content of Hsc Chemistry 1st Paper Question 2014 Out
  - Distinguishing Credible Sources

- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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