

Ojt Hrm

First week in my OJT

May 25 Friday

Our on-the job training in Tempura Japanese Restaurant took-off with a brief orientation with the store manager. First, the trainees were introduced to the whole staff, followed by a tour inside the store premises-from the dining area, then to the bar and finally in the kitchen. During this tour, the manager gave us a few ideas on how it works and we have our first glimpse on how the restaurant operates.

After that activity, he/she gave us our first on-site post: the entrance. We're instructed to stay there and open the door for the guests. He/She also reminded us that it's a must to greet the incoming guests with "Yokoso , Ma'am/ Sir" while outgoing customers must hear the words "Doumo arigatou gozaimasu".

May 26 Saturday

On our second day on Tempura, the staff taught us the standard position for serving. At the end of this activity, we were expected to learn the basic steps on how to serve the guests efficiently, have an adequate proficiency about the menu including the right pronunciation for each dish. A staff also made a demonstration on table-setting. In this task, the demonstrator specifically pointed-out that the table was set in a specific manner that follows a strict standard. After the activity, we were instructed to go back to our first designated post and usher the guests.

May 27 Sunday

After two days of warming-up on our new environment, the trainees were able to catch up on the life in Tempura. Day 3 is an exciting day because the manager allowed us to move more freely and gave us new set of assignments. First, they let us clean the tables. Aside from that, he/she instructed us to memorize the table count. This is important because today some of us were assigned in ushering the guests and serving the menu.

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Karin Nielsen-Saines



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for the nitration 4 01 ml of nitric acid 4 00 ml of sulfuric acid and 4 56 ml bromobenzene were used to complete the reaction with these measurements bromobenzene was found to be the limiting reagent since only 4 56 ml of bromobenzene was used the amount of product predicted to be recovered was 8 7533 g

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the aim of the experiment is to determine and study electrophilic aromatic substitution reaction eas particularly nitration of methyl benzoate go give methyl m nitrobenzoate thus determine the weight melting point and percentage yield of the pure methyl m nitrobenzoate abstract

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in this experiment we will carry out the nitration of bromobenzene bromine substituted on an aromatic ring is an ortho para directing group it will direct an incoming electrophile to the ortho and para positions it is also deactivating meaning that its reactions will be slower than the identical reactions of benzene

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in the nitration of bromobenzene a mixture of sulfuric acid and nitric acid react with one another nitric acid the weaker of the two acids will serve as base to form nitronium ion bromobenzene will attack the nitronium ion by nucleophilic attack

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then actual is your actual findings yield actual theoretical $\times 100$ what are the possible side reactions of nitration of bromobenzene dinitration can occur at ortho para or at both ortho positions occurs in overheating meta product happens very rarely because bromine is an ortho para director

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the bromine substituent in bromobenzene is an ortho para director the para position is favored due to steric effects in the reaction that takes place in this lab a nitro group is substituted onto bromobenzene at the para position to form 1-bromo-4-nitrobenzene

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study with quizlet and memorize flashcards containing terms like reacting sulfuric and nitric acid together produces the nitronium ion alone nitric acid is not a strong enough electrophile to react with bromobenzene sulfuric acid is added to the nitric acid the reverse rxn can overheat and boil easily includes a claisen adaptor and

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over 5 minutes slowly pipet this bromobenzene into the icy cold nitration mixture HNO_3 3 h 2 so 4 in the 125 ml flask and stir continue to stir the reaction mixture in the ice bath for 30 minutes do not heat in a new 250 ml erlenmeyer

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on the right the arenium cation has a nitro set with incredibly small 0.6 angstrom bond distances this will result in an optimization failure as the atoms are even close for the programme to accommodate their electrons demonstration of creating the p arenium cation by nitration of bromobenzene

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the purpose of this lab was to observe where substituents are added on a benzene ring by way of electrophilic aromatic substitution the electrophilic source in the nitration of bromobenzene comes from a nitronium ion NO_2^+ which acts as the electrophilic species

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nitration of bromobenzene is an electrophilic aromatic substitution reaction the electrophile will be formed from 4 0ml of nitric acid and 4 0ml of sulfuric acid and will then react with 4 5ml of bromobenzene this will produce o bromonitrobenzene and p bromonitrobenzene

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