

Hewlett-Packard Professional Certification

Examination 1017 has six to eight following problems. Use all appropriate and proper scientific units. Use scientific notation for answers requiring simplification of the number 1000.

1. Find the standard enthalpy change for the reaction $\text{C}_{10}\text{H}_8 + 12\text{H}_2\text{O} \rightarrow 10\text{H}_2\text{O} + 8\text{H}_2$
 of element data $\text{C}_{10}\text{H}_8 + 12\text{H}_2\text{O} \rightarrow 10\text{H}_2\text{O} + 8\text{H}_2$ $\Delta H^\circ = -1000 \text{ kJ}$
 and $\text{C}_{10}\text{H}_8 + 12\text{H}_2\text{O} \rightarrow 10\text{H}_2\text{O} + 8\text{H}_2$ $\Delta H^\circ = -1000 \text{ kJ}$

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2. Find standard enthalpy change for the formation of element and compound from
 $\text{C}_{10}\text{H}_8 + 12\text{H}_2\text{O} \rightarrow 10\text{H}_2\text{O} + 8\text{H}_2$ $\Delta H^\circ = -1000 \text{ kJ}$
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3. Find Gibbs free energy for the reaction of element and compound
 and hydrogen, which will give the equation $\text{C}_{10}\text{H}_8 + 12\text{H}_2\text{O} \rightarrow 10\text{H}_2\text{O} + 8\text{H}_2$ $\Delta H^\circ = -1000 \text{ kJ}$

Examine the standard enthalpy change for the reaction $\text{C}_{10}\text{H}_8 + 12\text{H}_2\text{O} \rightarrow 10\text{H}_2\text{O} + 8\text{H}_2$
 standard enthalpy $\text{C}_{10}\text{H}_8 + 12\text{H}_2\text{O} \rightarrow 10\text{H}_2\text{O} + 8\text{H}_2$ $\Delta H^\circ = -1000 \text{ kJ}$

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Padhraic Smyth



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