

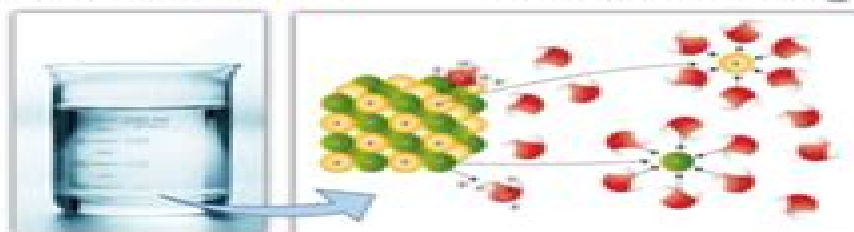
Chem102 Preparation of Solutions

HANDOUT 1

Solutions are homogenous mixtures that contains two or more substances. As we all know, solutions are homogenous combination of two or more compounds. When a solution is homogenous, the difference between the physical property of both solutions is cannot be distinguished. A solution has uniform composition. Examples are sugar and table salt in water. The dissolving medium (water) is the solvent and the dissolved substances (sugar or salt) is the solute.

SOLUTE **NaCl**

SOLVENT **H₂O**



Components:

1. Solvent
 - present in the greatest quantity
 - determines the phase of solution
 - if water is a component, it is usually the solvent.
2. Solute
 - component dissolved in the solvent

If the SOLUTE SOLVENT interaction is stronger than the solute-solute interaction and solvent solvent interaction then the solution process would be spontaneous. Water is a universal solvent but not all substances can dissolved by water.

Solubility is a temperature dependent

- maximum amount of solute that dissolves in a given amount of solution or solvent at a given temperature
- unit = grams of solute per mL of solvent

E.g. Solubility of NaCl = 35.7 g/100 mL at 20 °C

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