



Part III: Balance the following chemical equations

20.  $\underline{\hspace{1cm}} \text{Al} + \underline{\hspace{1cm}} \text{Cl}_2 \rightarrow \underline{\hspace{1cm}} \text{AlCl}_3$
21.  $\underline{\hspace{1cm}} \text{Al} + \underline{\hspace{1cm}} \text{FeO} \rightarrow \underline{\hspace{1cm}} \text{Al}_2\text{O}_3 + \underline{\hspace{1cm}} \text{Fe}$
22.  $\underline{\hspace{1cm}} \text{Na}_2\text{CO}_3 + \underline{\hspace{1cm}} \text{HCl} \rightarrow \underline{\hspace{1cm}} \text{H}_2\text{CO}_3 + \underline{\hspace{1cm}} \text{NaCl}$
23.  $\underline{\hspace{1cm}} \text{P}_4\text{O}_{10} + \underline{\hspace{1cm}} \text{H}_2\text{O} \rightarrow \underline{\hspace{1cm}} \text{H}_3\text{PO}_4$
24.  $\underline{\hspace{1cm}} \text{Al}_2(\text{SO}_4)_3 + \underline{\hspace{1cm}} \text{Ca}(\text{OH})_2 \rightarrow \underline{\hspace{1cm}} \text{Al}(\text{OH})_3 + \underline{\hspace{1cm}} \text{CaSO}_4$
25.  $\underline{\hspace{1cm}} \text{C}_3\text{H}_8 + \underline{\hspace{1cm}} \text{O}_2 \rightarrow \underline{\hspace{1cm}} \text{CO}_2 + \underline{\hspace{1cm}} \text{H}_2\text{O}$
26.  $\underline{\hspace{1cm}} \text{Fe}_2\text{O}_3 + \underline{\hspace{1cm}} \text{H}_2 \rightarrow \underline{\hspace{1cm}} \text{Fe} + \underline{\hspace{1cm}} \text{H}_2\text{O}$
27.  $\underline{\hspace{1cm}} \text{N}_2\text{O}_5 + \underline{\hspace{1cm}} \text{H}_2\text{O} \rightarrow \underline{\hspace{1cm}} \text{HNO}_3$
28.  $\underline{\hspace{1cm}} \text{H}_3\text{AsO}_4 \rightarrow \underline{\hspace{1cm}} \text{As}_2\text{O}_5 + \underline{\hspace{1cm}} \text{H}_2\text{O}$
29.  $\underline{\hspace{1cm}} \text{Pb} + \underline{\hspace{1cm}} \text{Na} + \underline{\hspace{1cm}} \text{C}_2\text{H}_5\text{Cl} \rightarrow \underline{\hspace{1cm}} \text{Pb}(\text{C}_2\text{H}_5)_4 + \underline{\hspace{1cm}} \text{NaCl}$

**Part IV:** In each of the following problems you are given some potential reaction. Determine what type of reaction would occur, predict the products, write the complete chemical equation and balance it!

30. Aluminum powder burns in the presence of oxygen to form ...
31. Magnesium metal reacts with nitric acid to form ...
32. Aqueous barium chloride reacts with sodium hydroxide to form ...
33. Butane burns in the presence of oxygen to form ...
34. Hydrogen gas reacts with chlorine gas to form ...
35. Aluminum powder reacts with iron(III) oxide to form ...

# Formulas And Chemical Reactions Study Guide

**R Barnett**



## **Formulas And Chemical Reactions Study Guide:**

This book delves into Formulas And Chemical Reactions Study Guide. Formulas And Chemical Reactions Study Guide is a vital topic that needs to be grasped by everyone, ranging from students and scholars to the general public. This book will furnish comprehensive and in-depth insights into Formulas And Chemical Reactions Study Guide, encompassing both the fundamentals and more intricate discussions.

1. This book is structured into several chapters, namely:
    - Chapter 1: Introduction to Formulas And Chemical Reactions Study Guide
    - Chapter 2: Essential Elements of Formulas And Chemical Reactions Study Guide
    - Chapter 3: Formulas And Chemical Reactions Study Guide in Everyday Life
    - Chapter 4: Formulas And Chemical Reactions Study Guide in Specific Contexts
    - Chapter 5: Conclusion
  2. In chapter 1, this book will provide an overview of Formulas And Chemical Reactions Study Guide. This chapter will explore what Formulas And Chemical Reactions Study Guide is, why Formulas And Chemical Reactions Study Guide is vital, and how to effectively learn about Formulas And Chemical Reactions Study Guide.
  3. In chapter 2, the author will delve into the foundational concepts of Formulas And Chemical Reactions Study Guide. This chapter will elucidate the essential principles that need to be understood to grasp Formulas And Chemical Reactions Study Guide in its entirety.
  4. In chapter 3, this book will examine the practical applications of Formulas And Chemical Reactions Study Guide in daily life. This chapter will showcase real-world examples of how Formulas And Chemical Reactions Study Guide can be effectively utilized in everyday scenarios.
  5. In chapter 4, this book will scrutinize the relevance of Formulas And Chemical Reactions Study Guide in specific contexts. This chapter will explore how Formulas And Chemical Reactions Study Guide is applied in specialized fields, such as education, business, and technology.
  6. In chapter 5, the author will draw a conclusion about Formulas And Chemical Reactions Study Guide. This chapter will summarize the key points that have been discussed throughout the book.
- This book is crafted in an easy-to-understand language and is complemented by engaging illustrations. It is highly recommended for anyone seeking to gain a comprehensive understanding of Formulas And Chemical Reactions Study Guide.

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