



Centre Number	Candidate Number

MINISTRY OF EDUCATION AND HUMAN RESOURCE DEVELOPMENT

SOLOMON ISLANDS YEAR 9 EXAMINATION

2020

ENGLISH

QUESTION AND ANSWER BOOKLET

TUESDAY 10TH NOVEMBER 9:00 AM

TIME: 2 HOURS plus 15
Minutes reading time

<u>QUESTION</u>	<u>CONTENT</u>	<u>MARKS</u>
1	Listening Comprehension	10
2	Reading Comprehension 1	10
3	Reading Comprehension 2	20
4	Fill in blank spaces	10
5	Grammar	30
6	Written Composition	20
TOTAL MARKS		100

INSTRUCTIONS TO CANDIDATES

1. Do not open this booklet until you are told to do so.
2. Write both your Centre Number and Candidate Number in the box provided at the top right hand corner of this page and at the end of this Booklet.
3. Before you begin each part, read through the instructions carefully.
4. Write your best answer letters to questions 1, 2 and 5 in the boxes provided at the end of this booklet.
5. Write your answers to questions 3, 4 and 6 in the spaces provided in this booklet.
6. Do NOT use correction fluid.
7. Mobile phones are NOT allowed in the examination room.
8. Page 18 is left blank deliberately.

YOU MUST HAND THIS BOOKLET TO THE SUPERVISOR AT THE END OF THE EXAMINATION.

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web sep 22 2021 the equilibrium system can be written in the general form $ce_{aq} + h_{aq} \rightleftharpoons ce_{aq} + h_{aq}$ label 7 the equilibrium constant expression for this reaction is $K = \frac{ce_{aq} \cdot h_{aq}}{ce_{aq} \cdot h_{aq}}$ label 8 where we denote the equilibrium constant K with a subscript a for acid

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web mar 25 2020 this activity supports the following unit topics and learning objectives unit 7 equilibrium topic 7 1 introduction to equilibrium tra 6 a explain the relationship between the occurrence of a reversible chemical or physical process and the establishment of equilibrium to experimental observations

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web write down the equilibrium equation you are investigating using the information provided by the introduction within the virtual lab activity equilibrium reaction $Co + H_2O \rightleftharpoons CoO + H_2$ select each control test tube to compare the concentrations of

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web the study of reaction rates will lead you into the study of chemical equilibrium in this semester you will also study electrochemistry from batteries to electroplating and you will learn about acids bases and salts unit 15 is a brief study of the atomic nucleus which your teacher may introduce at any time during the year

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web equilibrium moles $HSCN$ $2.00 \times 10^{-5} \text{ mol}$ $3.00 \times 10^{-6} \text{ mol}$ $1.70 \times 10^{-5} \text{ mol}$ $HSCN$ knowing the number of moles of Fe^{3+} and

hscn present in the equilibrium mixture and the volume of the mixture we can easily find the concentrations of those two species
 $\text{mol Fe}^{3+} = 1.70 \times 10^{-5} \text{ mol}$
 $\text{mol Fe}^{3+} = 8.50 \times 10^{-4} \text{ mol}$

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web jan 29 2023 7.1 introduction to equilibrium let's start this unit out with a broad overview of what equilibrium is all about in chemistry we define equilibrium as the point at which the rate of the forward reaction is equal to the rate of the reverse reaction we learned about rates in unit five as how quickly a reaction proceeds

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web equilibrium questions a covalent bond has a standard gibbs free energy of 400 kJ/mol while the hydrogen bond and van der Waals forces have substantially lower values respectively 25 and 5 kJ/mol given such values which of the following is equivalent to the ratio of the K_{eq} value of a covalent bond to that of a hydrogen bond

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web nov 13 2022 define the equilibrium state of a chemical reaction system what is its practical significance state the meaning and significance of the following terms reversible reaction quantitative reaction kinetically inhibited reaction explain the meaning of the statement equilibrium is macroscopically static but microscopically dynamic very

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web in a chemical equilibrium the forward and reverse reactions occur at equal rates and the concentrations of products and

reactants remain constant if we run a reaction in a closed system so that the products cannot escape we often find the reaction does not give a 100 yield of products