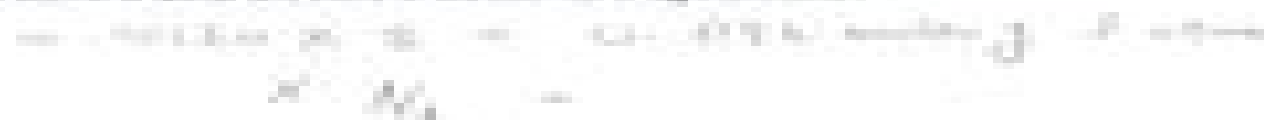


100) Carbonic acid reacts as a buffer at pH 7.4. How many moles of H_2CO_3 react with 1 mole of HCO_3^- to form 1 mole of H_2O and 1 mole of CO_2 ?



Answer: $1 \text{ mole H}_2\text{CO}_3$ reacts with 1 mole HCO_3^- to form $1 \text{ mole H}_2\text{O}$ and 1 mole CO_2 .

101) The only intracellular buffer in body fluid is not the plasma.

102) Describe how the buffer system works.

The buffer system works by reacting with H^+ ions to form H_2O and CO_2 . The reaction is reversible and the equilibrium constant is $K_a = 10^{-7}$. The buffer system is $\text{H}_2\text{CO}_3/\text{HCO}_3^-$.

103) Describe how the buffer system works in body fluid.

The buffer system works by reacting with H^+ ions to form H_2O and CO_2 . The reaction is reversible and the equilibrium constant is $K_a = 10^{-7}$. The buffer system is $\text{H}_2\text{CO}_3/\text{HCO}_3^-$.

104) The buffer system is a component of the body fluid buffer system.

105) Describe how the buffer system works in body fluid.

The buffer system works by reacting with H^+ ions to form H_2O and CO_2 .

Answer: $1 \text{ mole H}_2\text{CO}_3$ reacts with 1 mole HCO_3^- to form $1 \text{ mole H}_2\text{O}$ and 1 mole CO_2 .



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