

Pairs of electrons	BP	LP	Shape	Bond angle	E.g.
2	2	0	Linear	180°	CO ₂
3	3	0	<u>Trigonal planar</u>	120°	BF ₃
4	4	0	Tetrahedral	109.5°	CH ₄
4	3	1	Pyramidal	107°	NH ₃
4	2	2	Non-linear	104.5°	H ₂ O
6	6	0	Octahedral	90°	SF ₆

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