

Gas Laws Worksheet

$$\text{atm} = 760.0 \text{ mm Hg} = 101.3 \text{ kPa} = 760.0 \text{ torr}$$

Boyle's Law Problems: $P_1 V_1 = P_2 V_2$

1. If 22.5 L of nitrogen at 748 mm Hg are compressed to 725 mm Hg at constant temperature. What is the new volume?

$$(748 \text{ mmHg})(22.5 \text{ L}) = (725 \text{ mmHg}) V_2$$

$$V_2 = \frac{(748 \text{ mmHg})(22.5 \text{ L})}{(725 \text{ mmHg})}$$

$$V_2 = 23.2 \text{ L}$$

2. A gas with a volume of 4.0 L at a pressure of 205 kPa is allowed to expand to a volume of 12.0 L. What is the pressure in the container if the temperature remains constant?

$$(4.0 \text{ L})(205 \text{ kPa}) = (12.0 \text{ L}) P_2$$

$$P_2 = \frac{(4.0 \text{ L})(205 \text{ kPa})}{12.0 \text{ L}}$$

$$P_2 = 68.3 \text{ kPa}$$

3. What pressure is required to compress 196.0 liters of air at 1.00 atmosphere into a cylinder whose volume is 26.0 liters?

$$(196.0 \text{ L})(1.00 \text{ atm}) = (26.0 \text{ L}) P_2$$

$$P_2 = \frac{(196.0 \text{ L})(1.00 \text{ atm})}{26.0 \text{ L}}$$

$$P_2 = 7.54 \text{ atm}$$

4. A 40.0 L tank of ammonia has a pressure of 12.7 kPa. Calculate the volume of the ammonia if its pressure is changed to 8.4 kPa while its temperature remains constant.

$$(40.0 \text{ L})(12.7 \text{ kPa}) = (8.4 \text{ kPa}) V_2$$

$$V_2 = \frac{(40.0 \text{ L})(12.7 \text{ kPa})}{8.4 \text{ kPa}}$$

$$V_2 = 60.5 \text{ L}$$

Exploring Gas Laws Worksheet Answers

M Walker



Exploring Gas Laws Worksheet Answers:

Unveiling the Power of Verbal Art: An Emotional Sojourn through **Exploring Gas Laws Worksheet Answers**

In a global inundated with displays and the cacophony of instant communication, the profound power and emotional resonance of verbal art frequently fade into obscurity, eclipsed by the regular barrage of noise and distractions. However, set within the lyrical pages of **Exploring Gas Laws Worksheet Answers**, a captivating work of literary elegance that impulses with natural thoughts, lies an unforgettable trip waiting to be embarked upon. Penned with a virtuoso wordsmith, that mesmerizing opus courses visitors on an emotional odyssey, softly revealing the latent possible and profound influence embedded within the intricate web of language. Within the heart-wrenching expanse of this evocative evaluation, we shall embark upon an introspective exploration of the book's key styles, dissect its fascinating publishing model, and immerse ourselves in the indelible impression it leaves upon the depths of readers' souls.

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