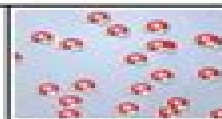


Activity B:

Temperature and molecular motion

Get the Gizmo ready:

- Click **Reset**, and select the **Micro** view.
- Set **Ice volume** to 0 cm³.
- Set **Add/remove heat energy** to 0 J/s.



Question: Why do phase changes occur?

1. **Compare:** Set the **Water temperature** to 0 °C and click **Play**. Observe the water molecules. Click **Reset**, set the **Water temperature** to 100 °C, and click **Play** again.

What do you notice?

Nothing differs, there is no difference.

2. **Observe:** Click **Reset**. The mean molecular speed of the water molecules is displayed below the container. Set the **Water temperature** to 0 °C and **Add/remove heat energy** to 400 J/s. Click **Play**.

A. How does the mean speed of the water molecules change as they are heated?

Increases the hotter water gets

B. Does the mean molecular speed change as much as the temperature as the water heats up? Explain.

Yes, as the mean molecular speed varies as temperature rises.

3. **Explain:** How is temperature related to the motions of molecules?

The molecules move more quickly as the temperature rises.

4. **Observe:** Click **Reset**. Set the **Water temperature** to 20 °C and the **Ice volume** to 50 cm³. Set **Add/remove heat energy** to 0 J/s. Click **Play**. How do the molecules in the liquid interact with the molecules in the solid?

The bonds are broken when the molecules of the liquid collide with those of the solid.

5. **Observe:** Click **Reset**. Set the **Water temperature** to 100 °C and the **Ice volume** to 50 cm³. Click **Play**. How does this situation compare to the previous one?

Ice melts more quickly

6. **Propose a theory:** Based on what you have observed, explain why you think phase changes occur. If possible, discuss your theory with your classmates and teacher.

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