

Teacher Guide: Boyle's Law and Charles' Law



Learning Objectives

- observe the effects of pressure, volume, and temperature
- define pressure
- describe how the pressure relates to the volume of a gas at a constant temperature (Boyle's law)
- describe how the temperature relates to the volume of a gas under constant pressure (Charles' law)



Essential

Students will: define Boyle's Law, Charles' Law, and related units, pressure



Learning Objectives

Boyle's Law: Boyle's Law states that the pressure and volume of a gas are inversely proportional, so that if the volume of a gas changes, the pressure will change in the opposite direction.

Boyle's Law: Boyle's Law states that the pressure and volume of a gas are inversely proportional, so that if the volume of a gas changes, the pressure will change in the opposite direction.

The following equations describe Boyle's Law:

- Boyle's Law: Pressure and volume are inversely proportional (Boyle's Law)
- Boyle's Law: Pressure and volume are inversely proportional (Boyle's Law)



Suggested Student Response

1. The student will:
 - observe the effects of pressure, volume, and temperature

Boyle's Law: Boyle's Law states that the pressure and volume of a gas are inversely proportional, so that if the volume of a gas changes, the pressure will change in the opposite direction.

Boyle's Law: Boyle's Law states that the pressure and volume of a gas are inversely proportional, so that if the volume of a gas changes, the pressure will change in the opposite direction.



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Sebastian Brünink



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