

## Experiment 9

## MOLAR MASS of a VOLATILE LIQUID

## Data Sheet

MEASUREMENTS	
A. Mass of Flask and Aluminum Foil Cover only	g
B. Gas Temperature	$^{\circ}\text{C}$ = K
C. **Mass of Flask, Aluminum Foil Cover, and Unknown Sample (at end of experiment ; recondensed vapor)	g
D. **Gas Volume (volume contained in flask)	mL = L
E. **Gas Pressure	mm Hg = atm
F. Mass of Unknown Sample (recondensed vapor at end of experiment)	g
G. Number of Moles of the Unknown sample (at end of the experiment  ( E x D ) /  (0.0821 L x atm x mol <sup>-1</sup> x K <sup>-1</sup> x B )	mol
H. Molar Mass of Unknown Sample ( F / G )	g / mol
I. Unknown Number	
J. Comparing Line H to Table 9.1, the Unknown is probably	

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# Experiment 9 Molar Mass Volatile Liquid

**H Kauffman**



## **Experiment 9 Molar Mass Volatile Liquid:**

*Laboratory Manual for Principles of General Chemistry* Jo Allan Beran, 2010-11-01 This new edition of the Beran lab manual emphasizes chemical principles as well as techniques The manual helps students understand the timing and situations for the various techniques The Beran lab manual has long been a market leading lab manual for general chemistry Each experiment is presented with concise objectives a comprehensive list of techniques and detailed lab intros and step by step procedures

*Laboratory Manual for Principles of General Chemistry* J. A. Beran, Mark Lassiter, 2022-08-16 Laboratory Manual for Principles of General Chemistry 11th Edition covers two semesters of a general chemistry laboratory program The material focuses on the lab experiences that reinforce the concepts that not all experimental conclusions are the same and depend on identifying an appropriate experimental procedure selecting the proper apparatus employing the proper techniques systematically analyzing and interpreting the data and minimizing inherent variables As a result of good data a scientific and analytical conclusion is made which may or may not be right but is certainly consistent with the data Experiments write textbooks textbooks don't write experiments A student's scientific literacy grows when experiences and observations associated with the scientific method are encountered Further experimentation provides additional cause effect observations leading to an even better understanding of the experiment The 11th edition's experiments are informative and challenging while offering a solid foundation for technique safety and experimental procedure The reporting and analysis of the data and the pre and post lab questions focus on the intuitiveness of the experiment The experiments may accompany any general chemistry textbook and are compiled at the beginning of each curricular unit An Additional Notes column is included in each experiment's Report Sheet to provide a space for recording observations and data during the experiment Continued emphasis on handling data is supported by the Data Analysis section

**Chemistry** Sally Solomon, Susan Rutkowski, Charles Boritz, 2008-05-02 Chemistry An Everyday Approach to Chemical Investigation is intended to accompany any mainstream general chemistry course and consists of 27 experiments that can be completed using only chemicals found in consumer products The manual is an ideal resource for courses emphasizing green chemistry in which the use of hazardous materials is reduced or eliminated altogether Many of the experiments requiring simple equipment and glassware can be performed at remote sites providing laboratory experience for use with on line or long distance learning courses The advantages of using accessible materials in chemistry laboratory are considerable Students can reinforce lecture discussions while working with familiar materials For instructors assembling the chemicals required for a lab course can be accomplished with limited budgets and without access to a chemical company Problems with safety and waste disposal are significantly reduced

*Chemical Principles in the Laboratory* Emil J. Slowinski, 1996 Provides a series of experiments designed to teach students the available experimental methods the proper design of experiments and the interpretation of experimental results

*Experiments in Chemistry* Frank R. Milio, Clyde R. Metz, W. G. Nordulf, 1991-03 **Chemistry for the IB Diploma**

**Workbook with CD-ROM** Jacqueline Paris, 2017-04-06 Chemistry for the IB Diploma Second edition covers in full the requirements of the IB syllabus for Chemistry for first examination in 2016 This workbook is specifically for the IB Chemistry syllabus for examination from 2016 The Chemistry for the IB Diploma Workbook contains straightforward chapters that build learning in a gradual way first outlining key terms and then providing students with plenty of practice questions to apply their knowledge Each chapter concludes with exam style questions This structured approach reinforces learning and actively builds students confidence using key scientific skills handling data evaluating information and problem solving This helps empower students to become confident and independent learners Answers to all of the questions are on the CD ROM

**Laboratory Experiments for Introduction to Chemistry** Thomas R. Dickson, 1975      **Methods of Experimental Physics**, 1967      Lcg Ol Chemistry,      Chemistry in the Laboratory Jo A. Beran, 1996 This updated and revised chemistry manual provides a well rounded understanding of concepts in the general chemistry laboratory Utilising visual aids experiments and equipment are explained and results and their pertinence discussed      **Experimental Heat Transfer, Fluid Mechanics and Thermodynamics 1993** M.D. Kelleher, R.K. Shah, K.R. Sreenivasan, Y. Joshi, 2012-12-02 The papers contained in this volume reflect the ingenuity and originality of experimental work in the areas of fluid mechanics heat transfer and thermodynamics The contributors are drawn from 27 countries which indicates how well the worldwide scientific community is networked The papers cover a broad spectrum from the experimental investigation of complex fundamental physical phenomena to the study of practical devices and applications A uniform outline and method of presentation has been used for each paper      Physical Chemistry for the Life Sciences Peter William Atkins, R. George Ratcliffe, Julio de Paula, Mark Wormald, 2023 This text provides a balanced presentation of the concepts of physical chemistry and their applications to biology and biochemistry Written to straddle the worlds of physical chemistry and the life sciences it shows students how the tools of physical chemistry can elucidate biological questions      **Fast Scanning Calorimetry** Christoph Schick, Vincent Mathot, 2016-06-28 In the past decades the scan rate range of calorimeters has been extended tremendously at the high end from approximately 10 up to 10 000 000 C s and more The combination of various calorimeters and the newly developed Fast Scanning Calorimeters FSC now span 11 orders of magnitude by which many processes can be mimicked according to the time scale s of chemical and physical transitions occurring during cooling heating and isothermal stays in case heat is exchanged This not only opens new areas of research on polymers metals pharmaceuticals and all kinds of substances with respect to glass transition crystallization and melting phenomena it also enables in depth study of metastability and reorganization of samples on an 1 to 1000 ng scale In addition FSC will become a crucial tool for understanding and optimization of processing methods at high speeds like injection molding The book resembles the state of the art in Thermal Analysis Calorimetry and is an excellent starting point for both experts and newcomers in the field

Introduction to Chemistry Thomas R. Dickson, 1991 Based on feedback from students and professors alike this

introductory textbook has been revised to offer material in a different sequence and expanded end of chapter questions A major theme of the text is the introduction explanation and illustration of the problem solving methods of beginning chemistry Approaches to solutions chemical problems and the unit equation factor label or dimensional analysis methods are explained in detail with numerous examples Relevant analogies and special topics continue to reinforce introduce and illustrate chemical concepts     Atomic and Electron Physics ,1967-01-01 Atomic and Electron Physics     **Chemical News and Journal of Industrial Science** ,1860     **Journal of Research of the National Bureau of Standards** United States. National Bureau of Standards,1949     **Instructor's Guide for Introductory Chemistry in the Laboratory** James F. Hall,1996     *Laboratory Manual to Accompany Chemistry, [by] Stanley R. Radcliff, Marjorie H. Navidi* Arthur D. Baker,1990     **Nuclear Science Abstracts** ,1966

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