

Stoichiometry III – Theoretical Yield

Teacher

If we know what chemical reaction is occurring, we can use mole ratios to determine the limiting and excess reactants. We can do this by using the starting masses of each reactant to see how much of a product could be produced if all of the reactant was consumed. We then compare the number of moles produced by each reactant, and the smallest value is the **theoretical yield** – the amount of a product produced if all of the limiting reactant is consumed.

Practice Problems

1. Given the following reaction: $C_3H_8 + 5O_2 \rightarrow 3CO_2 + 4H_2O$

a) Determine the limiting and excess reagents if you started with 14.8 g of C_3H_8 and 33.44 g of O_2 .

$$\frac{14.8 \text{ g } C_3H_8}{44.09 \text{ g } C_3H_8} \times \frac{1 \text{ mol } C_3H_8}{1 \text{ mol } C_3H_8} \times \frac{4 \text{ mol } H_2O}{1 \text{ mol } C_3H_8} \times \frac{18.02 \text{ g } H_2O}{1 \text{ mol } H_2O} = 24.2 \text{ g } H_2O$$

$$\frac{33.44 \text{ g } O_2}{32.00 \text{ g } O_2} \times \frac{1 \text{ mol } O_2}{1 \text{ mol } O_2} \times \frac{4 \text{ mol } H_2O}{5 \text{ mol } O_2} \times \frac{18.02 \text{ g } H_2O}{1 \text{ mol } H_2O} = 15.1 \text{ g } H_2O$$

Limiting: O_2
Excess: C_3H_8

b) What is the theoretical yield of H_2O . $15.1 \text{ g } H_2O$

2. Given the following reaction: $Al_2(SO_4)_3 + 6NaOH \rightarrow 3Na_2SO_4 + 2Al(OH)_3$

a) Determine the limiting and excess reagents if 8.1 g of $Al_2(SO_4)_3$ is reacted with 8.1 g of $NaOH$

$$\frac{8.1 \text{ g } Al_2(SO_4)_3}{294.14 \text{ g } Al_2(SO_4)_3} \times \frac{1 \text{ mol } Al_2(SO_4)_3}{1 \text{ mol } Al_2(SO_4)_3} \times \frac{2 \text{ mol } Al(OH)_3}{1 \text{ mol } Al_2(SO_4)_3} \times \frac{78.00 \text{ g } Al(OH)_3}{1 \text{ mol } Al(OH)_3} = 4.3 \text{ g } Al(OH)_3$$

$$\frac{8.1 \text{ g } NaOH}{40.00 \text{ g } NaOH} \times \frac{1 \text{ mol } NaOH}{1 \text{ mol } NaOH} \times \frac{2 \text{ mol } Al(OH)_3}{6 \text{ mol } NaOH} \times \frac{78.00 \text{ g } Al(OH)_3}{1 \text{ mol } Al(OH)_3} = 5.3 \text{ g } Al(OH)_3$$

Limiting: $Al_2(SO_4)_3$
Excess: $NaOH$

b) What is the theoretical yield of aluminum hydroxide. $4.3 \text{ g } Al(OH)_3$

3. Given the following reaction: $4Al_2O_3 + 9Fe \rightarrow 3Fe_3O_4 + 8Al$

a) Determine the limiting and excess reagents if 25.4 g of Al_2O_3 is reacted with 10.2 g of Fe .

$$\frac{25.4 \text{ g } Al_2O_3}{101.96 \text{ g } Al_2O_3} \times \frac{1 \text{ mol } Al_2O_3}{1 \text{ mol } Al_2O_3} \times \frac{3 \text{ mol } Fe_3O_4}{4 \text{ mol } Al_2O_3} \times \frac{231.55 \text{ g } Fe_3O_4}{1 \text{ mol } Fe_3O_4} = 43.3 \text{ g } Fe_3O_4$$

$$\frac{10.2 \text{ g } Fe}{55.85 \text{ g } Fe} \times \frac{1 \text{ mol } Fe}{1 \text{ mol } Fe} \times \frac{3 \text{ mol } Fe_3O_4}{9 \text{ mol } Fe} \times \frac{231.55 \text{ g } Fe_3O_4}{1 \text{ mol } Fe_3O_4} = 14.1 \text{ g } Fe_3O_4$$

Limiting: Fe
Excess: Al_2O_3

b) What is the theoretical yield of iron(II,III) oxide (aka magnetite).

$$14.1 \text{ g } Fe_3O_4$$

Experiment 3stoichiometry Answers

**Dennis W. Smithenry,Joan Gallager-
Bolos**



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The Word and the World Michael Mulkay, 2024-02-09 First published in 1985 The Word and the World is a significant empirically based contribution to the sociological analysis of scientists discourse and scientific culture It also offers a radical departure from established forms of sociological discourse which has far reaching implications for all areas of sociological study The central aim of this book is to widen the range of textual forms used in empirical sociological research and thereby to increase sociologists ability to furnish interesting valid and practically useful conclusions The unusual textual forms employed for the purposes of sociological analysis are the play the investigative inquiry the participant analyst dialogue and the parody One of the main substantive topics that provides the material for these explorations in form is the scientific debate This book will be of interest to students of sociology rhetoric linguistics and science

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Chemistry Harold R. Hunt,Toby F. Block,George M. McKelvy,2002 This established manual focuses on using non hazardous materials to teach the experimental nature of general chemistry Experiments are written to address students of various academic backgrounds and differing interests and abilities in chemistry While most experiments can be conducted in a single three hour period some have been designed to be completed over an extended time to illustrate that chemical systems do not work at an arbitrary schedule Suggestions are provided for combining experiments of shorter length and similar pedagogy

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Announcement University of Arkansas. Division of University Extension, 1913

Teaching General Chemistry Mansoor Niaz, 2008 The main objective of this monograph is to incorporate history and philosophy of science in the chemistry curriculum in order to provide students an overview of the dynamics of scientific research which involves controversies conflicts and rivalries among scientists that is the humanising aspects of science A major thesis of this book is the parallel between the construction of knowledge by the students and the scientists In looking for this relationship it is not necessary that ontogeny recapitulate phylogeny but rather to establish that students can face similar difficulties in conceptualising problems as those faced by the scientists in the past Given the vast amount of literature on students alternative conceptions misconceptions in science it is plausible to suggest that these can be considered not as mistakes but rather as tentative models leading to greater conceptual understanding Just as scientists resist changes in the hard core of their beliefs by offering auxiliary hypotheses students may adopt similar strategies Conceptual change in science education can thus be conceptualised as building of tentative models that provide greater explanatory power to students understanding

Cambridge IGCSE® Physical Science Chemistry Workbook Richard Harwood, Ian Lodge, 2017-02-16 Cambridge IGCSE Physical Science resources tailored to the 0652 syllabus for first examination in 2019 and all components of the series are endorsed by Cambridge International Examinations This Chemistry Workbook is tailored to the Cambridge IGCSE Physical Science 0652 syllabus for first examination in 2019 and is endorsed for learner support by Cambridge International Examinations The workbook covers both the Core and the Supplement material Developing students scientific skills the workbook exercises are complemented by self assessment checklists to help students evaluate their work as they go Answers are provided at the back of the book

Chemistry Neil D. Jespersen, Alison Hyslop, 2021-11-02 Chemistry The Molecular Nature of Matter 8th Edition continues to focus on the intimate relationship that exists between structure at the atomic molecular level and the observable macroscopic properties of matter Key revisions in this edition focus on three areas The deliberate inclusion of more updated real world examples that relate common real world student experiences to the science of chemistry Simultaneously examples and questions have been updated to align them with career concepts relevant to the environmental engineering biological pharmaceutical and medical sciences Providing students with transferable skills with a

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Pathobiology of Cardiovascular Injury H.L. Stone, William B. Weglicki, 2012-12-06 This volume represents a part of the scientific proceedings of the Sixth Annual Meeting of the American Section of the International Society for Heart Research that was held in Oklahoma City from September 13 15 1984 The chapters have been grouped according to the content of invited symposia on cardiovascular topics ranging from electrophysiology autonomic control of the circulation aging of the myocardium mechanisms of cardiovascular injury cultured heart cell studies etc Abstracts of all the papers presented at this meeting are included in Volume 16 Supp 1 of the Journal of Molecular and Cellular Cardiology Since our understanding of the pathobiology of cardiovascular injury requires integrating knowledge from both basic and clinical sciences the information presented in this volume extends from fundamental biochemical processes to pharmacological agents required to treat complex cardiac arrhythmias At the molecular level the structural lipids of the sarcolemma may be altered during myocardial ischemia the role of free radicals in this process is one emerging area of active study At the subcellular level plasma membranes provide the anatomical basis for maintenance of optimal potentials of excitable cells the anatomy and physiology of the sarcolemma ultimately determines the response to drugs and metabolic perturbants At the cellular level much progress has been made in our understanding of cultured and adult myocytes during ischemia and hypoxia in particular the electrophysiology of these model systems is well understood and several chapters address this approach

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