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Stoichiometry and Percent Yield Practice

Answer the following questions about theoretical yield and percent yield. Make sure you show all work, units, and round for sig figs.

1. Use the following equation to answer the questions below.



If I start this reaction with 57.0 grams of nitrogen with excess hydrogen, how many grams of ammonia (NH_3) can theoretically be produced?

$$57.0 \text{ g N}_2 \times \frac{1 \text{ mol N}_2}{28.014 \text{ g N}_2} \times \frac{2 \text{ mol NH}_3}{1 \text{ mol N}_2} \times \frac{17.031 \text{ g NH}_3}{1 \text{ mol NH}_3} = \boxed{69.3 \text{ g NH}_3}$$

If 60.0 grams of ammonia is actually produced, what was my percent yield of ammonia?

$$\frac{60.0}{69.3} \times 100 = \boxed{86.6\%}$$

2. Use the following equation to answer the questions below.



If 11.3 grams of carbon dioxide gas is used in this reaction, how many grams of lithium carbonate can theoretically be produced?

$$11.3 \text{ g CO}_2 \times \frac{1 \text{ mol CO}_2}{44.009 \text{ g CO}_2} \times \frac{1 \text{ mol Li}_2\text{CO}_3}{1 \text{ mol CO}_2} \times \frac{73.87 \text{ g Li}_2\text{CO}_3}{1 \text{ mol Li}_2\text{CO}_3} = \boxed{19.0 \text{ g Li}_2\text{CO}_3}$$

If 14.4 grams of lithium carbonate are actually produced, what was my percent yield?

$$\frac{14.4}{19.0} \times 100 = \boxed{75.8\%}$$

Experiment 7 Reaction Stoichiometry Percent Yield Answers

E Durkheim



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Table of Contents Experiment 7 Reaction Stoichiometry Percent Yield Answers

1. Understanding the eBook Experiment 7 Reaction Stoichiometry Percent Yield Answers
 - The Rise of Digital Reading Experiment 7 Reaction Stoichiometry Percent Yield Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Experiment 7 Reaction Stoichiometry Percent Yield Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Experiment 7 Reaction Stoichiometry Percent Yield Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Experiment 7 Reaction Stoichiometry Percent Yield Answers
 - Personalized Recommendations
 - Experiment 7 Reaction Stoichiometry Percent Yield Answers User Reviews and Ratings
 - Experiment 7 Reaction Stoichiometry Percent Yield Answers and Bestseller Lists
5. Accessing Experiment 7 Reaction Stoichiometry Percent Yield Answers Free and Paid eBooks
 - Experiment 7 Reaction Stoichiometry Percent Yield Answers Public Domain eBooks
 - Experiment 7 Reaction Stoichiometry Percent Yield Answers eBook Subscription Services
 - Experiment 7 Reaction Stoichiometry Percent Yield Answers Budget-Friendly Options
6. Navigating Experiment 7 Reaction Stoichiometry Percent Yield Answers eBook Formats

- ePub, PDF, MOBI, and More
 - Experiment 7 Reaction Stoichiometry Percent Yield Answers Compatibility with Devices
 - Experiment 7 Reaction Stoichiometry Percent Yield Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Experiment 7 Reaction Stoichiometry Percent Yield Answers
 - Highlighting and Note-Taking Experiment 7 Reaction Stoichiometry Percent Yield Answers
 - Interactive Elements Experiment 7 Reaction Stoichiometry Percent Yield Answers
 8. Staying Engaged with Experiment 7 Reaction Stoichiometry Percent Yield Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Experiment 7 Reaction Stoichiometry Percent Yield Answers
 9. Balancing eBooks and Physical Books Experiment 7 Reaction Stoichiometry Percent Yield Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Experiment 7 Reaction Stoichiometry Percent Yield Answers
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Experiment 7 Reaction Stoichiometry Percent Yield Answers
 - Setting Reading Goals Experiment 7 Reaction Stoichiometry Percent Yield Answers
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Experiment 7 Reaction Stoichiometry Percent Yield Answers
 - Fact-Checking eBook Content of Experiment 7 Reaction Stoichiometry Percent Yield Answers
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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