

# Experiment 34 Prelaboratory Assignment

## An Equilibrium Constant

Date \_\_\_\_\_ Lab Sec. \_\_\_\_\_ Name \_\_\_\_\_ Desk No. \_\_\_\_\_

1. Three parameters affect the absorbance of a sample. Which one is the focus of this experiment?

(1) Concentration of the solution

(2) Tube length

(3) Wavelength

Concentration is being focused on

2. Experimental Procedure, Part A.1, Table 34.1. A 3.00-ml. aliquot of 0.001 M NaSCN is diluted to 25.0 ml. with 0.2 M  $\text{Fe}(\text{NO}_3)_3$  and 0.1 M  $\text{HNO}_3$ .

a. How many moles of  $\text{SCN}^-$  are present?

$$\text{no. of moles of } \text{SCN}^- = \frac{0.001 \text{ M} \times 3 \text{ ml}}{1000 \text{ ml}} = 3 \times 10^{-6} \text{ mol}$$

- b. If all of the  $\text{SCN}^-$  is complexed with  $\text{Fe}^{3+}$  to form  $\text{FeNCS}^{2+}$ , what is the molar concentration of  $\text{FeNCS}^{2+}$ ?

$$3 \times 10^{-6} \text{ mol of } \text{SCN}^- \text{ produce } 3 \times 10^{-6} \text{ mol of } \text{FeNCS}^{2+}$$

$$\text{Total volume of diluted solution} = \frac{3 \text{ ml} + 22 \text{ ml}}{1000 \text{ ml}} = 25 \text{ ml}$$

3. Experimental Procedure, Part A.1. For preparing a set of standard solutions of  $\text{FeNCS}^{2+}$ , the equilibrium molar concentration of  $\text{FeNCS}^{2+}$  is assumed to equal the initial molar concentration of the  $\text{SCN}^-$  in the reaction mixture. Why is this assumption valid?

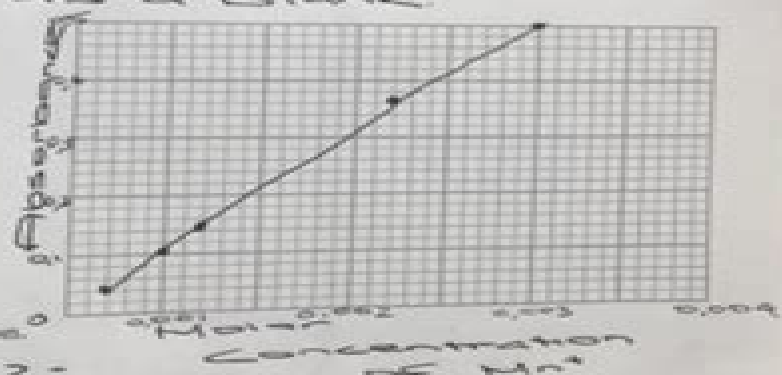
The  $\text{SCN}^-$  is the limiting reagent and thus when all of  $\text{SCN}^-$  is added its consumed no more complex formed. Concentration of  $\text{SCN}^-$  equals to concentration of  $\text{FeNCS}^{2+}$ .

4. Experimental Procedure, Part A.3. The blank solution used to calibrate the spectrophotometer is 10.0 ml. of 0.2 M  $\text{Fe}(\text{NO}_3)_3$  diluted to 25.0 ml. with 0.1 M  $\text{HNO}_3$ . Why is this solution preferred to simply using de-ionized water for the calibration?

The blank prepared above also takes into account the absorbance which must be subtracted to the correct value for the complex form when the actual reaction is done. So, plain water cannot be used as a blank.

5. Plot the following data as absorbance versus  $[\text{M}^{3+}]$  as a calibration curve:

Absorbance, A	Molar Concentration of $\text{M}^{3+}$
0.045	$3.0 \times 10^{-4} \text{ mol/L}$
0.097	$6.2 \times 10^{-4} \text{ mol/L}$
0.14	$9.0 \times 10^{-4} \text{ mol/L}$
0.35	$2.2 \times 10^{-3} \text{ mol/L}$
0.51	$3.2 \times 10^{-3} \text{ mol/L}$



A test solution showed a percent transmittance (%T) reading of 38.4%T. Interpret the calibration curve to determine the molar concentration of  $\text{M}^{3+}$  in the test solution.

$$\text{Absorbance} = 2 - \log \%T = 2 - \log(38.4) = 0.416$$

$$\text{Concentration of unknown solution (Mn}^{3+}) = 0.416 \times 3.2 \times 10^{-3} / 0.51 = 2.61 \times 10^{-3} \text{ mol/L}$$

# Experiment 34 An Equilibrium Constant Answers

**Joacim Rocklöv**



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