

Empirical and Molecular Formulas Practice

Read the question carefully for the data provided. Answers must be shown for credit.

Find empirical formula:

1. A compound is found to contain 40.0 g of carbon and 6.67 g of hydrogen. Find the empirical formula.

$$\text{C}_{40.0} \text{H}_{6.67} \times \frac{1 \text{ mole}}{12.01 \text{ g C}} = 3.33 \text{ moles C}$$

$$\text{H}_{6.67} \times \frac{1 \text{ mole}}{1.008 \text{ g H}} = 6.62 \text{ moles H}$$

C_{3.33}H_{6.62}

2. A compound is found to contain 50.0 g of carbon, 4.31 g of hydrogen, and 10.67 g of oxygen. Find the empirical formula.

$$\text{C}_{50.0} \text{H}_{4.31} \text{O}_{10.67} \times \frac{1 \text{ mole}}{12.01 \text{ g C}} = 4.16 \text{ moles C}$$

$$\text{H}_{4.31} \times \frac{1 \text{ mole}}{1.008 \text{ g H}} = 4.28 \text{ moles H}$$



$$\text{C}_{4.16} \text{H}_{4.28} \text{O}_{0.667} \times \frac{1 \text{ mole}}{0.667 \text{ moles O}} = \text{C}_{6.22} \text{H}_{6.41} \text{O}_1 = \text{C}_6\text{H}_6\text{O}$$

3. A 10.0 g sample of a compound, a gas at room temperature, is found to contain 8.56 g of carbon, 1.44 g of hydrogen, and 0.00 g of oxygen. Find the empirical formula.

$$\text{C}_{8.56} \text{H}_{1.44} \times \frac{1 \text{ mole}}{12.01 \text{ g C}} = 0.713 \text{ moles C}$$



$$\text{C}_{0.713} \text{H}_{1.43} \times \frac{1 \text{ mole}}{0.713 \text{ moles C}} = \text{C}_1 \text{H}_2 \text{O}_0 = \text{CH}_2$$

$$\text{C}_{0.713} \text{H}_{1.43} \times \frac{1 \text{ mole}}{1.008 \text{ g H}} = 1.43 \text{ moles H} \quad \text{C}_1 \text{H}_2 \text{O}_0 = \text{CH}_2$$

4. A 25.00 g sample of an unknown compound contains 15.38 g carbon, 4.09 g hydrogen, and 5.53 g oxygen. Find the empirical formula.

$$\text{C}_{15.38} \text{H}_{4.09} \text{O}_{5.53} \times \frac{1 \text{ mole}}{12.01 \text{ g C}} = 1.28 \text{ moles C}$$



$$\text{H}_{4.05} \times \frac{1 \text{ mole}}{1.008 \text{ g H}} = 4.05 \text{ moles H}$$

$$\text{O}_{5.53} \times \frac{1 \text{ mole}}{16.00 \text{ g O}} = 0.346 \text{ moles O} \quad \text{C}_{3.7} \text{H}_{11.7} \text{O}_1 = \text{C}_4\text{H}_{12}\text{O}$$

Empirical Formula Pogil Answers

M Mark



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