

KEY

Naming Ionic Compounds

What are the structural units that make up ionic compounds and how are they named?

What?

When working in chemistry, it is often convenient to write a chemical in symbols. For example we might write down the substance table salt as NaCl . In talking about chemistry however, it is a bit tacky to say "en-ay see-el" when we want to refer to a substance. Also, in formal writing we should use the name of the compound rather than its symbols. Therefore we need to learn how to say the proper names of ionic substances.

Model 1 – Ion Charges for Selected Elements

1	H ⁺									
2	Li ⁺	Be ²⁺						N ³⁻	O ²⁻	F ⁻
3	Na ⁺	Mg ²⁺	Transition elements				Al ³⁺	P ³⁻	S ²⁻	Cl ⁻
4	K ⁺	Ca ²⁺	Fe ²⁺ Fe ³⁺	Ni ²⁺ Ni ³⁺	Cu ⁺ Cu ²⁺	Zn ²⁺				Br ⁻
5	Rb ⁺	Sr ²⁺			Ag ¹⁺			Sn ²⁺ Sn ⁴⁺		I ⁻
6		Ba ²⁺				Hg ²⁺ Hg ¹⁺		Pb ²⁺ Pb ⁴⁺		
	Cations						Anions			

1. Based on the information in Model 1:
- Identify three elements that form only one cation.
 H^+ Li^+ Na^+ K^+ Rb^+ Ba^{+2} Mg^{+2} Ca^{+2} Sr^{+2} Ba^{+2} Zn^{+2} Ag^+ Al^{+3}
 - Identify three elements that form only one anion.
 N^{-3} P^{-3} O^{-2} S^{-2} F^- Cl^- Br^- I^-
 - Identify three elements that form more than one cation.
 Fe Ni Cu Hg Sn Pb
 - In what region of the periodic table are these "multiple ion" elements usually located?
Transition Metals

2. Consider the ions of potassium (K) and sulfur (S). Write chemical formulas for all possible ionic compounds involving these ions, using the simplest ratio(x) of potassium (K) and sulfur (S). Keep in mind that the sum of the charges in an ionic compound must equal zero.



3. Consider the ions of iron (Fe) and sulfur (S). Write chemical formulas for all possible ionic compounds involving these ions, using the simplest ratio(s) of iron (Fe) and sulfur (S). Keep in mind that the sum of the charges in an ionic compound must equal zero.



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Baby Professor



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Table of Contents Naming Compounds Pogil With Answers

1. Understanding the eBook Naming Compounds Pogil With Answers
 - The Rise of Digital Reading Naming Compounds Pogil With Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Naming Compounds Pogil With Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Naming Compounds Pogil With Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Naming Compounds Pogil With Answers
 - Personalized Recommendations
 - Naming Compounds Pogil With Answers User Reviews and Ratings
 - Naming Compounds Pogil With Answers and Bestseller Lists
5. Accessing Naming Compounds Pogil With Answers Free and Paid eBooks
 - Naming Compounds Pogil With Answers Public Domain eBooks
 - Naming Compounds Pogil With Answers eBook Subscription Services
 - Naming Compounds Pogil With Answers Budget-Friendly Options
6. Navigating Naming Compounds Pogil With Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Naming Compounds Pogil With Answers Compatibility with Devices
 - Naming Compounds Pogil With Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Naming Compounds Pogil With Answers
 - Highlighting and Note-Taking Naming Compounds Pogil With Answers
 - Interactive Elements Naming Compounds Pogil With Answers
8. Staying Engaged with Naming Compounds Pogil With Answers

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Naming Compounds Pogil With Answers
- 9. Balancing eBooks and Physical Books Naming Compounds Pogil With Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Naming Compounds Pogil With Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Naming Compounds Pogil With Answers
 - Setting Reading Goals Naming Compounds Pogil With Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Naming Compounds Pogil With Answers
 - Fact-Checking eBook Content of Naming Compounds Pogil With Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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