

- (T) A student estimated quantitatively  $x$ g of  $(\text{NH}_4)_2\text{SO}_4$ . Excess NaOH C, containing  $1.00\text{dm}^3$  of  $2.05\text{mol dm}^{-3}$  was heated with the  $(\text{NH}_4)_2\text{SO}_4$  to expel all the  $\text{NH}_3$  gas.  $25.0\text{cm}^3$  portion of excess solution C after treatment with  $(\text{NH}_4)_2\text{SO}_4$  was titrated with  $0.195\text{mol dm}^{-3}$   $\text{H}_2\text{SO}_4$  using few drops of phenolphthalein as indicator. The table below was presented as follows:

Burette readings ( $\text{cm}^3$ )	1st	2nd	3rd
Final burette reading	25.05	23.55	31.05
Initial burette reading	1.50	2.70	10.80
Volume of $\text{H}_2\text{SO}_4$ used	20.55	20.85	20.25

**FASTANSWER TEAM**

The equation of the reaction between NaOH solution and  $\text{H}_2\text{SO}_4$  solution is:



[Na = 23.0; O = 16.0; H = 1.00; S = 32.0; N = 14.0]

- a. (i) Mention the laboratory apparatus used to measure the  $25.0\text{cm}^3$  portion of solution C and state why it should be used?

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- (ii) State whether phenolphthalein indicator is suitable or unsuitable. Give reason for the answer.

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- (iii) State why few drops of phenolphthalein indicator was used.

\_\_\_\_\_

- (iv) Explain briefly why the student needed to repeat another titration exercise.

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- (v) State three experimental errors that might be committed by the student.

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# Nabteb Question Paper

**Scott C. Dulebohn**



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