#### Ideal Gas Law Worksheet PV = nRT

Use the ideal gas law, "PerV-nRT", and the universal gas constant  $R = 0.0821 \frac{L*atm}{L*mol}$  to solve the following problems:

\*\*K\*mol\*\*

If pressure is needed in kPa then convert by multiplying by 101.3kPa / latm to get

R = 8.31 kPa\*L/(K\*mole)

1) If I have 4 moles of a gas at a pressure of 5.6 atm and a volume of 12 liters, what is the temperature?

2) If I have an unknown quantity of gas at a pressure of 1.2 atm, a volume of 31 liters, and a temperature of 87 °C, how many moles of gas do I have?

3) If I contain 3 moles of gas in a container with a volume of 60 liters and at a temperature of 400 K, what is the pressure inside the container?

If I have 7.7 moles of gas at a pressure of 0.09 atm and at a temperature of 56 °C,

# J Weston Walch Gas Law Problems Answers

**Padhraic Smyth** 

J Weston Walch Gas Law Problems Answers:

Whispering the Techniques of Language: An Psychological Quest through J Weston Walch Gas Law Problems Answers

In a digitally-driven earth where monitors reign supreme and instant transmission drowns out the subtleties of language, the profound secrets and psychological subtleties concealed within words usually go unheard. Yet, nestled within the pages of **J Weston Walch Gas Law Problems Answers** a interesting fictional value pulsating with fresh emotions, lies a fantastic journey waiting to be undertaken. Published by an experienced wordsmith, this charming opus invites viewers on an introspective trip, gently unraveling the veiled truths and profound impact resonating within the very cloth of every word. Within the emotional depths of this touching evaluation, we will embark upon a heartfelt exploration of the book is core styles, dissect their interesting writing type, and fail to the powerful resonance it evokes serious within the recesses of readers hearts.

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