

# চূড়ান্ত সাজেশন HSC 2023(Chemistry 1<sup>st</sup> Paper CQ)

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1st Paper, Chapter-2 : গুণগত রসায়ন( 2 সেট CQ আসবে)

Part-1: পরমানু, কোয়ান্টাম সংখ্যা, H এর পারমাণবিক বর্ণালি ( 1 সেট CQ থাকবে)

- ★1) বোরারফোর্ড পরমানু মডেল ও সীমাবদ্ধতা
- ★★2) বোর পরমানু মডেল ও সীমাবদ্ধতা ও  $mvr = nh/2\pi$
- ★★★3) পরমানুর তরঙ্গ অথবা ঔর্ধ কক্ষপথে কতগুলো অরবিটাল বা ইলেকট্রন থাকতে পারে কোয়ান্টাম সংখ্যার সাহায্যে বের করো(ছক)
- ★★★4) কক্ষপথের ইলেকট্রনের ধাপান্তরের ফলে সৃষ্ট বর্ণালির তরঙ্গ দৈর্ঘ্য অথবা বর্ণ অথবা শক্তি নির্ণয় করো।
- ★★5) সর্বশেষ ইলেকট্রনের জন্য পাউলির বর্জন নীতি দেখাও

Part-2: দ্রাব্যতা ও আয়ন শনাক্তকরণ(1 সেট CQ আসবে)

- ★★★1) দ্রাব্যতা দেওয়া থাকবে  $K_{sp}$  বের করো
- ★★★2)  $K_{sp}$  দেওয়া দ্রাব্যতা অথবা আয়নের ঘনমাত্রা বের করো
- ★★★3) দুটি দ্রবনকে মেশালে অধঃক্ষেপ পড়বে কিনা গাণিতিক ভাবে বিশ্লেষণ করো।
- ★★4) আয়ন শনাক্তকরণ : অম্লীয় :  $NH_4^+, Al^{3+}, Ca^{2+}, Na^+, Cu^{2+}$  (less important  $Fe^{2+}, Fe^{3+}$ )  
অম্লীয় :  $Cl^-, SO_4^{2-}, CO_3^{2-}$

Extra: ★★5) দ্রবন ঠাণ্ডা করলে কতটুকু অধঃক্ষেপ পড়বে

★★6) অপমাত্রা বাড়ালে অতিরিক্ত কতটুকু লবন দ্রবীভূত করলে দ্রবনটি সম্পৃক্ত হবে

★★7) সম আয়ন  $m$

NB: Part-2 থেকে 1, 3 প্রশ্নটি আসার সম্ভাবনা বেশি।

1st Paper, Chapter-3 : মৌলের পর্যায়বৃত্ত ধর্ম ও রাসায়নিক বন্ধন( 2 সেট CQ আসবে)

★★★1) আয়নীকরণ শক্তি:

★★★i) Be থেকে B এর আয়নীকরণ শক্তি কম কেন?

★★★ii) N থেকে O এর আয়নীকরণ শক্তি কম কেন?

★★★2) ইলেক্ট্রন আসক্তি:

★★★i) F থেকে Cl এর ইলেক্ট্রন আসক্তি বেশি কেন?

★★★3) সংকরন: প্রশ্ন হবে সংকরণ ব্যাখ্যা কর, গঠন, আকৃতি, বন্ধন কোন এত কেন অথবা তুলনা করো।

i) যাদের  $sp$  ইলেক্ট্রন নেই:  $CH_3PCl_2, SF_6, BF_3$  ★★  $CH_2=CH_2$  ★★  $CH \equiv CH$  or  $C_2H_2$

ii) যাদের  $sp$  ইলেক্ট্রন আছে: ★★  $NH_3$  ★★  $H_2O, PH_3, H_2S$

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