

- A. $\text{H}_2\text{O}(g) + 2\text{CO}(g) \rightarrow \text{heat} + 2\text{H}_2\text{O}(l)(g)$
- B. $\text{H}_2\text{O}(g) + 2\text{H}_2\text{O}(g) \rightarrow 2\text{H}_2\text{O}(g) + \text{heat}$
- C. $\text{CH}_4(g) + \text{H}_2\text{O}(g) \rightarrow \text{heat} + \text{CO}(g) + 2\text{H}_2\text{O}(g)$
- D. $\text{C}_2\text{H}_2(g) + 4\text{H}_2(g) \rightarrow \text{CH}_4(g) + 2\text{H}_2\text{S}(g) + \text{heat}$

26. Consider the following equilibrium system: $\text{H}_2(g) + 2\text{H}_2\text{O}(g) \rightleftharpoons 2\text{H}_2\text{O}(g) + 92 \text{ kJ}$
In order to maximize the yield for this reaction, the best conditions are:

- A. Low pressure and low temperature
- B. Low pressure and high temperature
- C. High temperature and low pressure
- D. High pressure and low temperature

27. Consider the following equilibrium: $\text{CaCO}_3(s) \rightleftharpoons 556 \text{ kJ} \rightleftharpoons \text{CaCO}_3(s) + \text{CO}_2(g)$
The equilibrium will shift to the right:

- A. CO_2 is added
- B. $\text{CaCO}_3(s)$ is added
- C. The temperature is increased
- D. The temperature is decreased.

28. Explain the difference between Arrhenius acids and bases and Brønsted-Lowry acids and bases.

29. List three characteristics that acids and bases have in common.

30. To raise the pH of a solution, would you add HCl or NaOH? Why?

31. Describe and give an example of an electrolyte.

32. Find the percent composition of each element in $\text{NH}_4\text{C}_2\text{H}_3\text{O}_2$.

33. Given the balanced equation: $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$, how many grams of H_2O can be produced from 50g of O_2 ?

34. Draw energy diagrams for the following:
Reactants 50kJ Products 20kJ
Products 30kJ Products 40kJ
Activation energy 20kJ Activation energy 50kJ

35. Given the unbalanced equation below, what mass of sodium chloride can be produced from 23 g of Na and 35g of Cl_2 ?



36. Explain the difference between saturated and unsaturated carbon compounds.

37. Identify the nuclear forces and explain each.

38. Explain how chemical reactions differ from nuclear reactions.

Honors Chemistry Final Exams

Robert R. Redfield



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