

CHEMISTRY I EQUATION SHEET

Solubility Rules

Solubility Rules			Form a Compound which is	
Negative Ion Plus Positive Ion				
Any neg. ion + group 1a ions Soluble				
Any neg. ion + Ammonium ion Soluble				
Nitrate + any positive ion Soluble				
Acetate	+	any positive ion except Ag ⁺ or Hg ²⁺	Soluble	
bromide or iodide + Any other positive ion	+	Ag ⁺ , Pb ²⁺ , Hg ²⁺ or Cu ⁺	Not Soluble	Soluble
Sulfate	+	Ca ²⁺ , Sr ²⁺ , Ba ²⁺ , Ra ²⁺ , Ag ⁺ or Pb ²⁺	Not Soluble	Soluble
+ Any other positive ion Soluble				
Sulfide	+	group 1a ions or ammonium	Soluble	
Be ²⁺ , Mg ²⁺ , Ca ²⁺ , Sr ²⁺ , Ba ²⁺ or Ra ²⁺			Soluble	
+ Any other positive ion Not Soluble				
Hydroxide	+	group 1a ions or ammonium	Soluble	
+ Any other positive ion Not Soluble				
Carbonate, Chromate, Phosphate, or Sulfite + Any other positive ion	+	group 1a ions or ammonium	Soluble	
Not Soluble				

Equations

D = m/v	c = specific heat
Molarity = $\frac{\text{mol}}{\text{V}}$	d = density
$P_1 V_1 = P_2 V_2$	I = initial amount
$\frac{T_1}{T_2}$	F = final amount
Q = mc ΔT	m = mass
nRTPV =	M = molarity
F = I (1/2) *	n = number of moles & number of half lives
n = Total Time	P = pressure
Half Life Time	R = ideal gas constant
	Q = heat energy
	T = temperature
	V = volume
	X = amount of matter left

Constants

	1 atm = 101.3 kPa Pressure
	1 atm = 760 mm Hg
	R = 8.314 L x kPa / mol°C Universal
Constant or 0.0821 L x atm / mol°C	
Avogadro's number	6.02 X
Specific Heats (J / g°C)	
	Water = 4.180
	Aluminum = 0.895
	Copper = 0.387
	Iron = 0.448

Activity Series in Decreasing Activity

Metals	
Lithium	
Rubidium	
Potassium	
Calcium	
Sodium	
Magnesium	
Aluminum	
Manganese	
Zinc	
Iron	
Nickel	
Tin	
Lead	
Hydrogen	
Copper	
Silver	
Platinum	
Gold	
Nonmetals	
Fluorine	
Chlorine	
Bromine	
Iodine	

Polyatomic Ions

Name	Formula
Ammonium	NH ₄ ⁺
Acetate	C ₂ H ₃ O ₂ ⁻
Carbonate	CO ₃ ⁻²
Bicarbonate or Hydrogen carbonate	HCO ₃ ⁻
Hypochlorite	ClO ⁻
Chlorite	ClO ₂ ⁻
Chlorate	ClO ₃ ⁻
Perchlorate	ClO ₄ ⁻
Hydroxide	OH ⁻
Cyanide	CN ⁻
Nitrate	NO ₃ ⁻
Nitrite	NO ₂ ⁻
Chromate	CrO ₄ ⁻²
Oxalate	C ₂ O ₄ ⁻²
Sulfate	SO ₄ ⁻²
Sulfite	SO ₃ ⁻²
Hydrogen sulfate	HSO ₄ ⁻
Phosphate	PO ₄ ⁻³
Phosphite	PO ₃ ⁻³
Hydrogen phosphate	HPO ₄ ⁻²

Metric Conversions

Unit	Abbr.	Meaning
Mega-	M	10 ⁶
kilo-	k	10 ³
hecto-	h	10 ²
deca-	da	10 ¹
Liter	L	1
gram	g	1
meter	m	1
deci-	d	10 ⁻¹
centi-	c	10 ⁻²
milli-	m	10 ⁻³
micro-	μ	10 ⁻⁶
nano-	n	10 ⁻⁹

Calorie Conversions

Water = 1 cal/ g°C
1000 cal = 1 Cal

High School Chemistry Formulas Cheat Sheet

Javier García-Martínez, Elena Serrano-Torregrosa



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