

CHAPTER 21 REVIEW

Nuclear Chemistry**SECTION 1****SHORT ANSWER** Answer the following questions in the space provided.

1. _____ Based on the information about the three elementary particles in the text, which has the greatest mass?
 - (a) the proton
 - (b) the neutron
 - (c) the electron
 - (d) They all have the same mass.
2. _____ The force that keeps nucleons together is
 - (a) a strong nuclear force.
 - (b) a weak nuclear force.
 - (c) an electromagnetic force.
 - (d) a gravitational force.
3. _____ The stability of a nucleus is most affected by the
 - (a) number of neutrons.
 - (b) number of protons.
 - (c) number of electrons.
 - (d) ratio of neutrons to protons.
4. _____ If an atom should form from its constituent particles,
 - (a) matter is lost and energy is taken in.
 - (b) matter is lost and energy is released.
 - (c) matter is gained and energy is taken in.
 - (d) matter is gained and energy is released.
5. _____ For atoms of a given mass number, those with greater mass defects, have
 - (a) smaller binding energies per nucleon.
 - (b) greater binding energies per nucleon.
 - (c) the same binding energies per nucleon as those with smaller mass defects.
 - (d) variable binding energies per nucleon.
6. Based on **Figure 1.1** of the text, which isotope of He, helium-3 or helium-4,
 - _____ a. has the smaller binding energy per nucleon?
 - _____ b. is more stable to nuclear changes?
7. The number of neutrons in an atom of magnesium-25 is _____.
8. Nuclides of the same element have the same _____.

Glencoe Chemistry Chapter 6 Answers

Alton Biggs



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Astronomy Alton Biggs, 2002 **Glencoe Algebra 1**, 2001 **Chemistry** McGraw-Hill Staff, 2001-03 **Electrons, Atoms, and Molecules in Inorganic Chemistry** Joseph J. Stephanos, Anthony W. Addison, 2017-06-01

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1. Understanding the eBook Glencoe Chemistry Chapter 6 Answers
 - The Rise of Digital Reading Glencoe Chemistry Chapter 6 Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Glencoe Chemistry Chapter 6 Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Glencoe Chemistry Chapter 6 Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Glencoe Chemistry Chapter 6 Answers
 - Personalized Recommendations
 - Glencoe Chemistry Chapter 6 Answers User Reviews and Ratings
 - Glencoe Chemistry Chapter 6 Answers and Bestseller Lists
5. Accessing Glencoe Chemistry Chapter 6 Answers Free and Paid eBooks
 - Glencoe Chemistry Chapter 6 Answers Public Domain eBooks
 - Glencoe Chemistry Chapter 6 Answers eBook Subscription Services
 - Glencoe Chemistry Chapter 6 Answers Budget-Friendly Options

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 - ePub, PDF, MOBI, and More
 - Glencoe Chemistry Chapter 6 Answers Compatibility with Devices
 - Glencoe Chemistry Chapter 6 Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Glencoe Chemistry Chapter 6 Answers
 - Highlighting and Note-Taking Glencoe Chemistry Chapter 6 Answers
 - Interactive Elements Glencoe Chemistry Chapter 6 Answers
8. Staying Engaged with Glencoe Chemistry Chapter 6 Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Glencoe Chemistry Chapter 6 Answers
9. Balancing eBooks and Physical Books Glencoe Chemistry Chapter 6 Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Glencoe Chemistry Chapter 6 Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Glencoe Chemistry Chapter 6 Answers
 - Setting Reading Goals Glencoe Chemistry Chapter 6 Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Glencoe Chemistry Chapter 6 Answers
 - Fact-Checking eBook Content of Glencoe Chemistry Chapter 6 Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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