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CHAPTER 12**STUDY GUIDE FOR CONTENT MASTERY****Section 12.3 Limiting Reactants**

In your textbook, read about why reactions stop and how to determine the limiting reactant.

Study the diagram showing a chemical reaction and the chemical equation that represents the reaction. Then complete the table. Show your calculations for questions 25–27 in the space below the table.



The molar masses of O_2 , N_2 , and NO_2 are 32.00 g/mol, 28.02 g/mol, and 46.01 g/mol, respectively.

Amount of O_2	Amount of N_2	Amount of NO_2	Limiting Reactant	Amount and Mass of Excess Reactant
1 molecule	2 molecules	2 molecules	none	none
2 molecules	2 molecules	4 molecules	N_2	2 molecules O_2
2 molecules	4 molecules	4 molecules	O_2	2 molecules N_2
1.00 mol	2.00 mol	2.00 mol	none	none
4.00 mol	2.00 mol	4.00 mol	N_2	2.00 mol O_2
6.00 mol	1.00 mol	2.00 mol	N_2	5.00 mol O_2
1.00 mol	4.00 mol	2.00 mol	O_2	3.00 mol N_2
0.500 mol	0.200 mol	0.400 mol	N_2	0.300 mol O_2
32.00 g	56.04 g	92.02 g	none	none
64.00 g	56.04 g	92.02 g	O_2	32.00 g N_2
64.00 g	112.08 g	92.02 g	O_2	48.04 g N_2

balanced equation mole ratio = 2 mol NO_2 /1 mol O_2

$16.00 \text{ g } \text{O}_2 \times 1 \text{ mol } \text{O}_2 / 32.00 \text{ g } \text{O}_2 = 0.500 \text{ mol } \text{O}_2$

$20.00 \text{ g } \text{N}_2 \times 1 \text{ mol } \text{N}_2 / 28.02 \text{ g } \text{N}_2 = 0.714 \text{ mol } \text{N}_2$

actual mole ratio = $0.714 \text{ mol } \text{N}_2 / 0.500 \text{ mol } \text{O}_2 = 1.428 \text{ mol } \text{N}_2 / 1.000 \text{ mol } \text{O}_2$

Because the actual mole ratio of N_2 to O_2 is larger than the balanced equation mole ratio of N_2 to O_2 , there is an excess of N_2 ; O_2 is the limiting reactant.

Mass of NO_2 used = $0.500 \text{ mol } \text{O}_2 \times 2 \text{ mol } \text{NO}_2 / 1 \text{ mol } \text{O}_2 = 1.000 \text{ mol } \text{NO}_2$

$1.000 \text{ mol } \text{NO}_2 \times 46.01 \text{ g } \text{NO}_2 / 1 \text{ mol } \text{NO}_2 = 46.01 \text{ g } \text{NO}_2$

Mass of NO_2 produced = $0.500 \text{ mol } \text{O}_2 \times 2 \text{ mol } \text{NO}_2 / 1 \text{ mol } \text{O}_2 = 1.000 \text{ mol } \text{NO}_2$

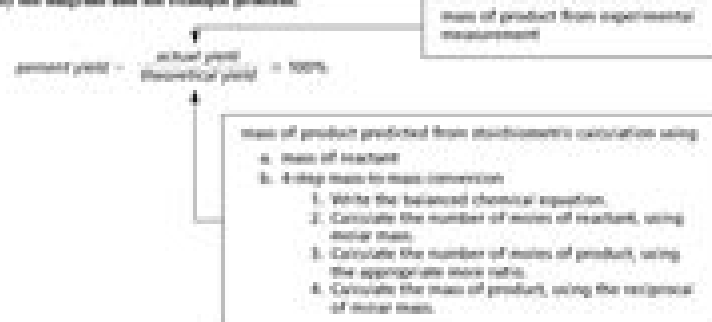
Excess N_2 = $20.00 \text{ g } \text{N}_2 - 16.00 \text{ g } \text{N}_2 = 4.00 \text{ g } \text{N}_2$

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CHAPTER 13**STUDY GUIDE FOR CONTENT MASTERY****Section 13.4 Percent Yield**

In your textbook, read about the yield of products.

Study the diagram and the example problem.



Example Problem: The following chemical equation represents the production of palladium, a valuable metal in the manufacturing of some semiconductor devices.



In one experiment, the reaction yielded 7.42 g of the metal from a 7.00-g sample of palladium. Determine the percent yield of this reaction. The molar masses of Ga and Ga_2O_3 are 69.72 g/mol and 157.44 g/mol, respectively.

Use the information in the diagram and example problem to evaluate each value or expression below. If the value or expression is correct, write correct. If it is incorrect, write the correct value or expression.

- actual yield: unknown 7.42 g Ga_2O_3
- mass of reactant: 7.00 g Ga correct
- number of moles of reactant: $7.00 \text{ g Ga} \times \frac{69.72 \text{ g Ga}}{1 \text{ mol Ga}} = 7.00 \text{ g Ga} \times 1 \text{ mol Ga} / 69.72 \text{ g Ga}$
- number of moles of product: $0.100 \text{ mol Ga} \times \frac{2 \text{ mol } (\text{Ga}_2\text{O}_3)}{1 \text{ mol Ga}} = 0.200 \text{ mol Ga} \times \frac{2 \text{ mol } (\text{Ga}_2\text{O}_3)}{1 \text{ mol Ga}}$
- theoretical yield: $0.200 \text{ mol } (\text{Ga}_2\text{O}_3) \times \frac{157.44 \text{ g } (\text{Ga}_2\text{O}_3)}{1 \text{ mol } (\text{Ga}_2\text{O}_3)} = 31.49 \text{ g } (\text{Ga}_2\text{O}_3)$ correct
- percent yield: $\frac{7.42 \text{ g } (\text{Ga}_2\text{O}_3)}{31.49 \text{ g } (\text{Ga}_2\text{O}_3)} \times 100 = 23.56\%$ correct

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