Directions: Match each term in Column I with its description in Column II. Write the letter of the correct term in the space provided.

# Column I binary compound chemically stable 3. nonpolar molecule \_\_\_\_\_ 4. ion \_\_\_\_ 5. ionic bond 6. polar molecule 7. oxidation number 8. chemical formula 9. covalent bond \_\_\_\_\_ 10. hydrate ----- 11. chemical bond \_\_\_\_\_ 12. polyatomic ion

# Column II

a. number that indicates how many electrons an atom must gain, lose, or share to become stable

Class

- b. shorthand that tells what elements a compound contains and the exact number of atoms of each element in a unit of the compound
- c. positively or negatively charged, covalently bonded group of atoms
- compound composed of two elements
- e. describes an atom that has a full outermost energy level
- f. molecule that has a slightly positive end and a slightly negative end
- g. the attraction that forms between atoms when they share electrons
- h. the force that holds atoms together in a compound
- a compound that has water chemically attached to it
- j. the force of attraction between the opposite charges of the ions in an ionic compound
- k. molecule made of two identical atoms that share the electrons equally
- a charged particle that has either more or fewer electrons than protons

# **Glencoe Science Physics Chapter 1 Assessment Answers**

**Minjie Lin** 

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Merrill Earth Science Ralph M. Feather, Susan Leach Snyder, Dale T. Hesser, 1995

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# **Table of Contents Glencoe Science Physics Chapter 1 Assessment Answers**

- 1. Understanding the eBook Glencoe Science Physics Chapter 1 Assessment Answers
  - The Rise of Digital Reading Glencoe Science Physics Chapter 1 Assessment Answers
  - Advantages of eBooks Over Traditional Books
- 2. Identifying Glencoe Science Physics Chapter 1 Assessment Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Glencoe Science Physics Chapter 1 Assessment Answers
  - User-Friendly Interface
- 4. Exploring eBook Recommendations from Glencoe Science Physics Chapter 1 Assessment Answers
  - Personalized Recommendations
  - Glencoe Science Physics Chapter 1 Assessment Answers User Reviews and Ratings

- Glencoe Science Physics Chapter 1 Assessment Answers and Bestseller Lists
- 5. Accessing Glencoe Science Physics Chapter 1 Assessment Answers Free and Paid eBooks
  - Glencoe Science Physics Chapter 1 Assessment Answers Public Domain eBooks
  - Glencoe Science Physics Chapter 1 Assessment Answers eBook Subscription Services
  - Glencoe Science Physics Chapter 1 Assessment Answers Budget-Friendly Options
- 6. Navigating Glencoe Science Physics Chapter 1 Assessment Answers eBook Formats
  - o ePub, PDF, MOBI, and More
  - Glencoe Science Physics Chapter 1 Assessment Answers Compatibility with Devices
  - Glencoe Science Physics Chapter 1 Assessment Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Glencoe Science Physics Chapter 1 Assessment Answers
  - Highlighting and Note-Taking Glencoe Science Physics Chapter 1 Assessment Answers
  - Interactive Elements Glencoe Science Physics Chapter 1 Assessment Answers
- 8. Staying Engaged with Glencoe Science Physics Chapter 1 Assessment Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Glencoe Science Physics Chapter 1 Assessment Answers
- 9. Balancing eBooks and Physical Books Glencoe Science Physics Chapter 1 Assessment Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Glencoe Science Physics Chapter 1 Assessment Answers
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Glencoe Science Physics Chapter 1 Assessment Answers
  - Setting Reading Goals Glencoe Science Physics Chapter 1 Assessment Answers
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Glencoe Science Physics Chapter 1 Assessment Answers
  - Fact-Checking eBook Content of Glencoe Science Physics Chapter 1 Assessment Answers
  - Distinguishing Credible Sources

- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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