

**Acids Nomenclature****Worksheet # 1- Naming Binary Acids**

You will now learn to name and write formulas for acids. The first type of acid we will name is a binary acid. Recall that binary compounds contain two different elements. The cation in all acids is the hydrogen ion or  $H^{+1}$ .

All binary acids are named using the following rule: hydro\_\_\_\_\_ic acid. The name of the element goes in the blank. For example:

HCl would be hydrochloric acid

H<sub>2</sub>S would be hydrosulfuric acid

Name the following acids.

Formula	Name	Formula	Name
HBr		H <sub>2</sub> Se	
HI		HF	
H <sub>3</sub> P		H <sub>2</sub> Te	

When you write the formulas for acids you use the same method you used when writing the formula for compounds containing metals.

- First you write each symbol with the charge. (Remember that the cation will always be  $H^{+1}$ )
- Next you switch the charges and make them subscripts. (Swap and Drop)

For example: hydroiodic acid  $\rightarrow H^{+1}$  and  $I^{-1}$  becomes HI

hydrophosphoric acid  $\rightarrow H^{+1}$  and  $P^{-3}$  becomes H<sub>3</sub>P

Name	Symbol w/ Oxidation Number	Formula
hydrosulfuric acid		
hydrochloric acid		
hydroarsenic acid		
hydrofluoric acid		
hydrosodic acid		
hydrotelluric acid		

# Naming Acids Packet

**James Knowles**

A red circular graphic with a gradient, appearing as a stylized arrow or a partial circle, located to the right of the author's name.

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